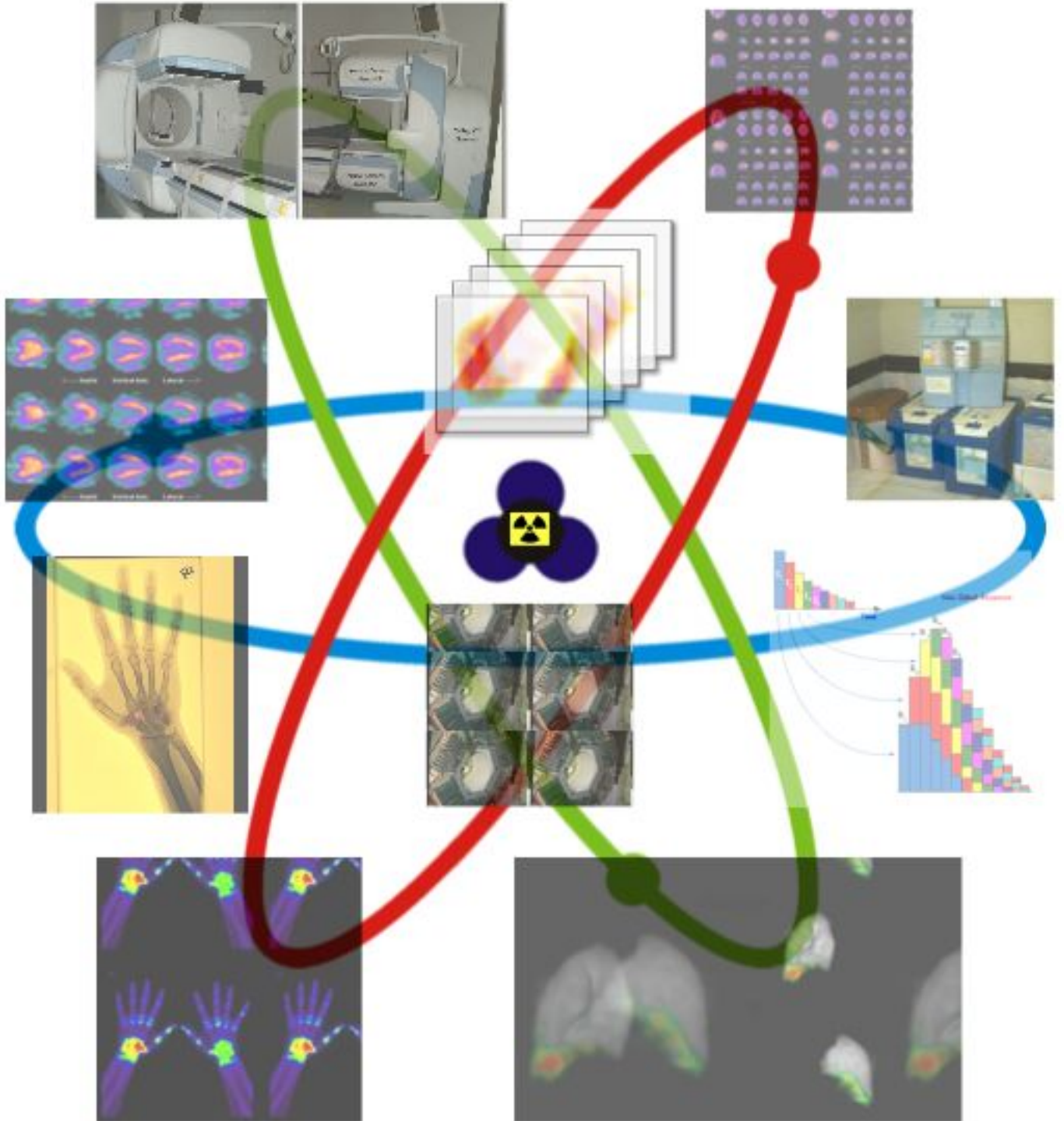


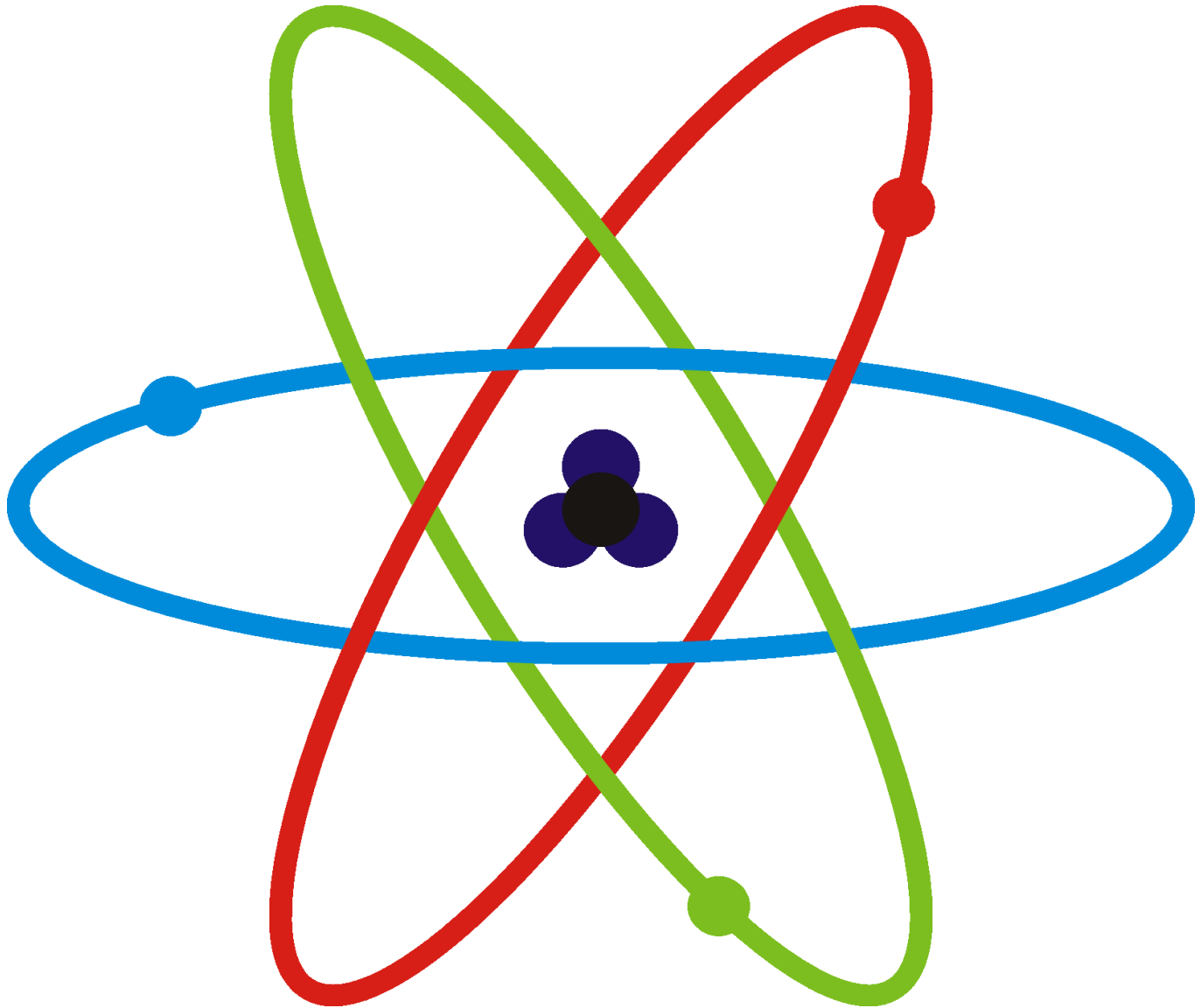
BASIC PHYSICS OF NUCLEAR MEDICINE

By Kieran Maher and other Wikibooks contributors.



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By *Kieran Maher* and other *Wikibooks* contributors.



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Contents

INTRODUCTION.....	3
EXTERNAL LINKS.....	3
CHAPTERS.....	4
01. Atomic & Nuclear Structure.....	4
02. Radioactive Decay.....	11
03. The Radioactive Decay Law.....	20
04. Units of Radiation Measurement.....	32
05. Interaction of Radiation with Matter.....	37
06. Attenuation of Gamma-Rays.....	42
07. Gas-Filled Radiation Detectors.....	55
08. Scintillation Detectors.....	63
09. Nuclear Medicine Imaging Systems.....	75
10. Production of Radioisotopes.....	90
CHAPTER REVIEW.....	94
ABOUT THE BOOK.....	101
History & Document Notes.....	101
Authors, Bibliography, & Image Credits.....	102
GNU Free Documentation License.....	103

Introduction

Nuclear Medicine is a fascinating application of **nuclear physics**. This wikibook is intended to support a basic introductory course in an early semester of an undergraduate program. It assumes that students have completed decent high school programs in maths and physics and are concurrently taking subjects in the medical sciences.

External Links

- **Australian & New Zealand Society of Nuclear Medicine** - the main nuclear medicine society in Australia; there is also a **Victorian Branch** website.
- **Radiation and Us** - covers common sources of natural and artificial radiation - in the environment, in the home, in the workplace and in medicine.
- **Society of Nuclear Medicine** - the main nuclear medicine society in the USA - describes what nuclear medicine is and how it relates to other medical imaging modalities. Career information is also included.

CHAPTER 1 ATOMIC & NUCLEAR STRUCTURE

Introduction

You will have encountered much of what we will cover here in your high school physics. We are going to review this material again below so as to set the context for subsequent chapters. This chapter will also provide you with an opportunity to check your understanding of this topic.

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The chapter covers atomic structure, nuclear structure, the classification of nuclei, binding energy and nuclear stability.

Atomic Structure

The **atom** is considered to be the basic building block of all matter. A simple theory of the atom tells us that it consists of two components: a **nucleus** surrounded by an **electron cloud**. The situation can be considered as being similar in some respects to planets orbiting the sun.

From an electrical point of view, the nucleus is said to be positively charged and the electrons negatively charged.

From a size point of view, the radius of an atom is about 10^{-10} m while the radius of a nucleus is about 10^{-14} m, i.e. about ten thousand times smaller. The situation could be viewed as something like a cricket ball, representing the nucleus, in the middle of a sporting arena with the electrons orbiting somewhere around where the spectators would sit. This perspective tells us that the atom should be composed mainly of empty space. However, the situation is far more complex than this simple picture portrays in that we must also take into account the physical forces which bind the atom together.

Chemical phenomena can be thought of as interactions between the electrons of individual atoms.

Radioactivity on the other hand can be thought of as changes which occur within the nuclei of atoms.

The Nucleus

A simple description of the nucleus tells us that it is composed of **protons** and **neutrons**. These two particle types are collectively called **nucleons**, i.e. particles which inhabit the nucleus.

From a mass point of view the mass of a proton is roughly equal to the mass of a neutron and each of these is about 2,000 times the mass of an electron. So most of the mass of an atom is concentrated in the small region at its core.

From an electrical point of view the proton is positively charged and the neutron has no charge. An atom all on its own (if that were possible to achieve!) is electrically neutral. The number of protons in

the nucleus of such an atom must therefore equal the number of electrons orbiting that atom.

Classification of Nuclei

The term **Atomic Number** is defined in nuclear physics as the number of protons in a nucleus and is given the symbol **Z**. From your chemistry you will remember that this number also defines the position of an element in the **Periodic Table of Elements**.

Click **HERE** to access an interactive webpage which gives details of elements in the Periodic Table.

The term **Mass Number** is defined as the number of nucleons in a nucleus, that is the number of protons plus the number of neutrons, and is given the symbol **A**.

Note that the symbols here are a bit odd, in that it would prevent some confusion if the Atomic Number were given the symbol A, and the Mass Number were given another symbol, such as M, but its not a simple world!

It is possible for nuclei of a given element to have the same number of protons but differing numbers of neutrons, that is to have the same Atomic Number but different Mass Numbers. Such nuclei are referred to as **Isotopes**. All elements have isotopes and the number ranges from three for hydrogen to over 30 for elements such as caesium and barium.

Chemistry has a relatively simple way of classifying the different elements by the use of symbols such as H for hydrogen, He for helium and so on. The classification scheme used to identify different isotopes is based on this approach with the use of a superscript before the chemical symbol to denote the Mass Number along with a subscript before the chemical symbol to denote the Atomic Number. In other words an isotope is identified as:

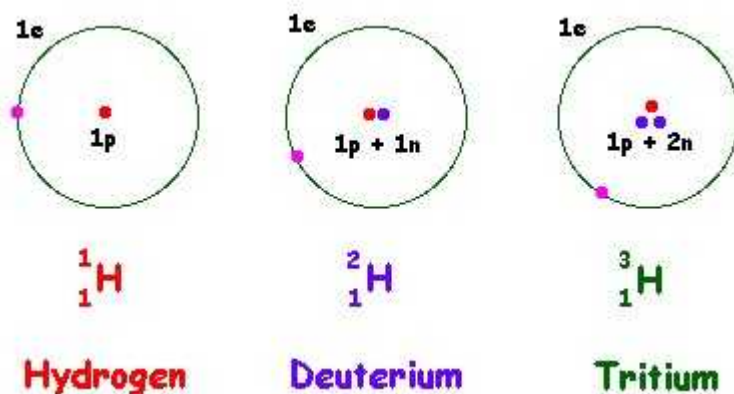


where X is the chemical symbol of the element.

Let us take the case of **hydrogen** as an example. It has three isotopes:

- the most common one consisting of a single proton orbited by one electron,
- a second isotope consisting of a nucleus containing a proton and a neutron orbited by one electron,
- a third whose nucleus consists of one proton and two neutrons, again orbited by a single electron.

A simple illustration of these isotopes is shown below. Remember though that this is a simplified illustration given what we noted earlier about the size of a nucleus compared with that of an atom. But the illustration is nevertheless useful for showing how isotopes are classified.



Click [HERE](#) for an animated version of this figure.

The first isotope commonly called **hydrogen** has a Mass Number of 1, an Atomic Number of 1 and hence is identified as:



The second isotope commonly called **deuterium** has a Mass Number of 2, an Atomic Number of 1 and is identified as:



The third isotope commonly called **tritium** is identified as:



The same classification scheme is used for all isotopes. For example, you should now be able to figure out that the uranium isotope, ${}^{236}_{92}\text{U}$, contains 92 protons and 144 neutrons.

A final point on classification is that we can also refer to individual isotopes by giving the name of the element followed by the Mass Number. For example, we can refer to deuterium as hydrogen-2 and we can refer to



as uranium-236.

Before we leave this classification scheme let us further consider the difference between chemistry and nuclear physics. You will remember that the water molecule is made up of two hydrogen atoms bonded with an oxygen atom. Theoretically if we were to combine atoms of hydrogen and oxygen in this manner many, many of billions of times we could make a glass of water. We could also make our glass of water using deuterium instead of hydrogen. This second glass of water would theoretically be very similar from a chemical perspective. However, from a physics perspective our second glass would

be heavier than the first since each deuterium nucleus is about twice the mass of each hydrogen nucleus. Indeed water made in this fashion is called **heavy water**.

Atomic Mass Unit

The conventional unit of mass, the **gram**, is rather large for use in describing characteristics of nuclei. For this reason, a special unit called the Atomic Mass Unit (amu) is often used. This unit is sometimes defined as 1/12th of the mass of the stable most commonly occurring isotope of carbon, i.e. ^{12}C . In terms of grams, 1 amu is equal to 1.66×10^{-24} g, that is, just over one million, million, million millionth of a gram.

The masses of the proton, m_p and neutron, m_n on this basis are:

$$m_p = 1.00783 \text{ amu}$$

and

$$m_n = 1.00866 \text{ amu}$$

while that of the electron is just 0.00055 amu.

Binding Energy

We are now in a position to consider the subject of nuclear stability. From what we have covered so far, we have seen that the nucleus is a tiny region in the centre of an atom and that it is composed of neutrally and positively charged particles. So, in a large nucleus such as that of uranium ($Z=92$) we have a large number of positively charged protons concentrated into a tiny region in the centre of the atom. An obvious question which arises is that with all these positive charges in close proximity, how come the nucleus does not fly apart? How can a nucleus remain as an entity with such electrostatic repulsion between the components? Should the orbiting negatively-charged electrons not attract the protons away from the nucleus?

Let us take the case of the **helium-4** nucleus as an example. This nucleus contains two protons and two neutrons so that in terms of amu we can figure out from what we covered earlier that the

$$\text{mass of 2 protons} = 2.01566 \text{ amu,}$$

and the

$$\text{mass of 2 neutrons} = 2.01732 \text{ amu.}$$

Therefore we would expect the total mass of the nucleus to be 4.03298 amu.

The experimentally determined mass of a helium-4 nucleus is a bit less - just 4.00260 amu. In other words there is a difference of 0.03038 amu between what we might expect as the mass of this

nucleus and what we actually measure. You might think of this difference as very small at just 0.75%. But remember that since the mass of one electron is 0.00055 amu the difference is actually equivalent to the mass of about 55 electrons. Therefore it is significant enough to wonder about.

It is possible to consider that this missing mass is converted to energy which is used to hold the nucleus together, that is it is converted to a form of energy called the **Binding Energy**. You could say like all relationships, energy must be expended in maintaining them!

Like the gram in terms of the mass of nuclei, the common unit of energy, the **joule** is rather cumbersome when we consider the energy needed to bind a nucleus together. The unit used to express energies on the atomic scale is the **electron volt**, symbol: eV.

One **electron volt** is defined as the amount of energy gained by an electron as it falls through a potential difference of one volt. This definition on its own is not of great help to us here and it is stated purely for the sake of completeness. So do not worry about it for the time being. Just appreciate that it is a unit representing a tiny amount of energy but an amount nevertheless which is useful on the atomic scale. It is a bit too small in the case of binding energies however and the mega-electron volt (MeV) is often used.

Albert Einstein introduced us to the equivalence of mass, m , and energy, E , at the atomic level using the following equation:

$$E = m c^2,$$

where c is the **velocity of light**.

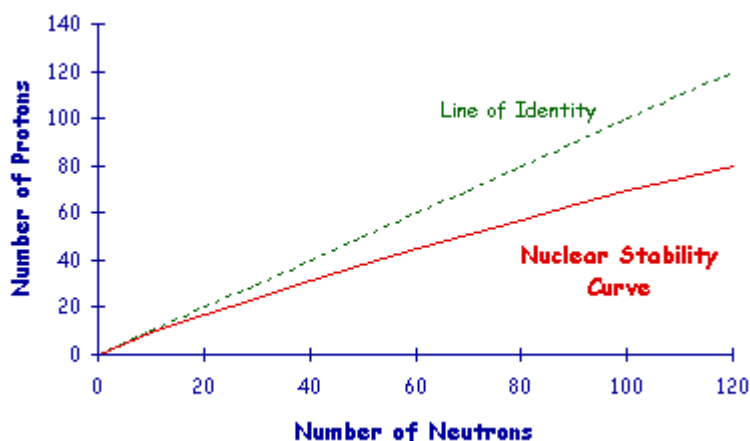
It is possible to show that 1 amu is equivalent to 931.48 MeV. Therefore, the mass difference we discussed earlier between the expected and measured mass of the helium-4 nucleus of 0.03038 amu is equivalent to about 28 MeV. This represents about 7 MeV for each of the four nucleons contained in the nucleus.

Nuclear Stability

In most stable isotopes the **binding energy** per nucleon lies between 7 and 9 MeV. Since this binding energy is dependent on the number of nucleons in a nucleus, that is the Mass Number, A , and since the electrostatic repulsion between the protons is dependent on the square of the charges, that is Z^2 , it is possible to infer that Z^2 must be dependent on A for a nucleus to remain stable.

In other words to combat the effect of the increase in electrostatic repulsion when the number of protons increases the number of neutrons must increase more rapidly to contribute sufficient energy to bind the nucleus together.

As we noted earlier there are a number of isotopes for each element of the Periodic Table. It has been found that the most stable isotope for each element has a specific number of neutrons in its nucleus. Plotting a graph of the number of protons against the number of neutrons for these stable isotopes generates what is called the **Nuclear Stability Curve**:



Note that the number of protons equals the number of neutrons for small nuclei. But notice also that the number of neutrons increases more rapidly than the number of protons as the size of the nucleus gets bigger so as to maintain the stability of the nucleus. In other words more neutrons need to be there to contribute to the binding energy used to counteract the electrostatic repulsion between the protons.

Radioactivity

There are about 2,450 known isotopes of the one hundred odd elements in the Periodic Table. You can imagine the size of a **table of isotopes** relative to that of the Periodic Table! The unstable isotopes lie above or below the Nuclear Stability Curve. These unstable isotopes attempt to reach the stability curve by splitting into fragments, in a process called **Fission**, or by emitting particles and/or energy in the form of radiation. This latter process is called **Radioactivity**.

It is useful to dwell for a few moments on the term radioactivity. For example what has nuclear stability to do with radio? From a historical perspective remember that when these radiations were discovered about 100 years ago we did not know exactly what we were dealing with. When people like **Henri Becquerel** and **Marie Curie** were working initially on these strange emanations from certain natural materials it was thought that the radiations were somehow related to another phenomenon which also was not well understood at the time - that of radio communication. It seems reasonable on this basis to appreciate that some people considered that the two phenomena were somehow related and hence that the materials which emitted radiation were termed *radio-active*.

We know today that the two phenomena are not directly related but we nevertheless hold onto the term radioactivity for historical purposes. But it should be quite clear to you having reached this stage of this chapter that the term radioactive refers to the emission of particles and/or energy from unstable isotopes. Unstable isotopes for instance those that have too many protons to remain a stable entity are called **radioactive isotopes** - and called **radioisotopes** for short. The term **radionuclide** is also sometimes used.

Finally about 300 of the 2,450-odd isotopes mentioned above are found in nature. The rest are man-made (or person-made!), that is they are produced artificially. These 2,150 or so artificial isotopes have been made during the last 100 years or so with most having been made since the second world war. Click **HERE** to access an interactive webpage which gives details of elements in the Periodic

Table and of radioisotopes used in medicine.

We will return to the production of radioisotopes in the last chapter of this wikibook and will proceed for the time being with a description of the types of radiation emitted by radioisotopes.

External Links

- [Marie and Pierre Curie and the Discovery of Polonium and Radium](#) - an historical essay from The Nobel Foundation.
- [Natural Radioactivity](#) - an overview of radioactivity in nature - includes sections on primordial radionuclides, cosmic radiation, human produced radionuclides, as well as natural radioactivity in soil, in the ocean, in the human body and in building materials - from the University of Michigan Student Chapter of the Health Physics Society.
- [The Particle Adventure](#) - an interactive tour of the inner workings of the atom which explains the modern tools physicists use to probe nuclear and sub-nuclear matter and how physicists measure the results of their experiments using detectors - from the Particle Data Group at the Lawrence Berkeley National Lab, USA and mirrored at CERN, Geneva.
- [WebElements](#) - an excellent web-based Periodic Table of the Elements which includes a vast array of data about each element - originally from Mark Winter at the University of Sheffield, England.

Multiple Choice Questions

Click [HERE](#) to access an online MCQ covering the material in this chapter.

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CHAPTER 01 • ATOMIC & NUCLEAR STRUCTURE

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CHAPTER 2 RADIOACTIVE DECAY

Introduction

We saw in the last chapter that radioactivity is a process used by unstable nuclei to achieve a more stable situation. It is said that such nuclei *decay* in an attempt to achieve stability. So, an alternative title for this chapter is **Nuclear Decay Processes**.

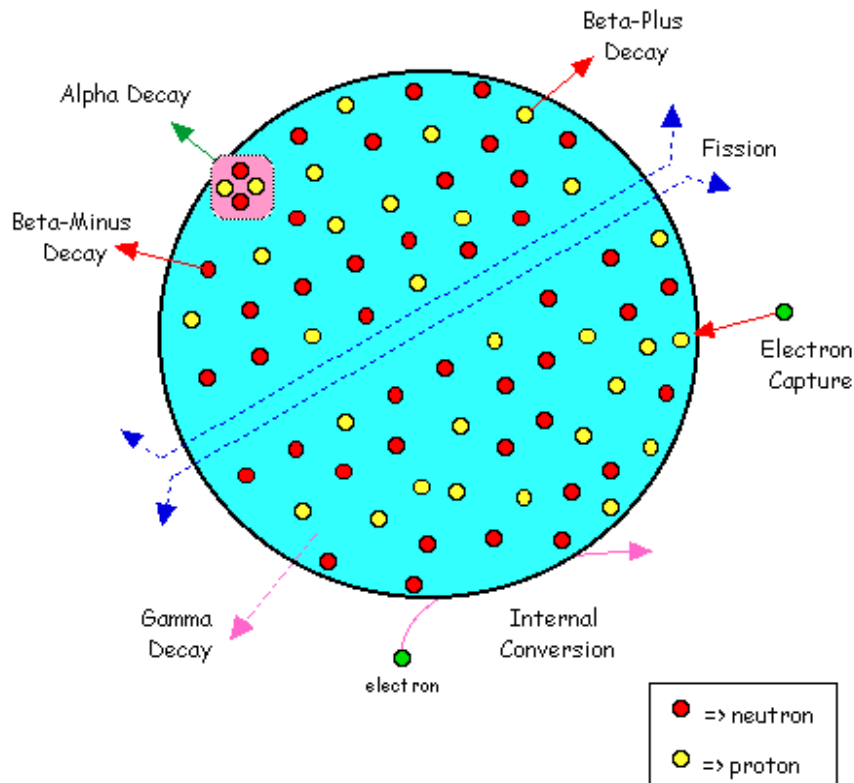
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We also saw in the previous chapter that we can use the Nuclear Stability Curve as a means of describing what is going on. So a second alternative title for this chapter is **Methods of Getting onto the Nuclear Stability Curve**.

We are going to follow a descriptive or phenomenological approach to the topic here by describing in a fairly simple fashion what is known about each of the major decay mechanisms. Once again you may have already covered this material in high school physics. But bear with us because the treatment here will help us set the scene for subsequent chapters.

Methods of Radioactive Decay

Rather than considering what happens to individual nuclei it is perhaps easier to consider a hypothetical nucleus that can undergo many of the major forms of radioactive decay. This hypothetical nucleus is shown below:



Click [HERE](#) for an animated version of this figure.

Firstly we can see two protons and two neutrons being emitted together in a process called **alpha-decay**. Secondly, we can see that a proton can release a particle in a process called **beta-plus decay**, and that a neutron can emit a particle in a process called **beta-minus decay**. We can also see an electron being captured by a proton. Thirdly we can see some energy being emitted which results from a process called **gamma-decay** as well as an electron being attracted into the nucleus and being ejected again. Finally there is the rather catastrophic process where the nucleus cracks in half called **spontaneous fission**. We will now describe each of these decay processes in turn.

Spontaneous Fission

This is a very destructive process which occurs in some heavy nuclei which split into 2 or 3 fragments plus some neutrons. These fragments form new nuclei which are usually radioactive. **Nuclear reactors** exploit this phenomenon for the production of radioisotopes. Its also used for nuclear power generation and in nuclear weaponry. The process is not of great interest to us here and we will say no more about it for the time being.

Alpha Decay

In this decay process two protons and two neutrons leave the nucleus together in an assembly known as an **alpha particle**. Note that an **alpha particle** is really a helium-4 nucleus.

So why not call it a helium nucleus? Why give it another name? The answer to this question lies in the history of the discovery of radioactivity. At the time when these radiations were discovered we

didn't know what they really were. We found out that one type of these radiations had a double positive charge and it was not until sometime later that we learnt that they were in fact nuclei of helium-4. In the initial period of their discovery this form of radiation was given the name alpha rays (and the other two were called beta and gamma rays), these terms being the first three letters of the Greek alphabet. We still call this form of radiation by the name **alpha particle** for historical purposes. Calling it by this name also contributes to the specific jargon of the field and leads outsiders to think that the subject is quite specialised!

But notice that the radiation really consists of a helium-4 nucleus emitted from an unstable larger nucleus. There is nothing strange about helium since it is quite an abundant element on our planet. So why is this radiation dangerous to humans? The answer to this question lies with the energy with which they are emitted and the fact that they are quite massive and have a double positive charge. So when they interact with living matter they can cause substantial destruction to molecules which they encounter in their attempt to slow down and to attract two electrons to become a neutral helium atom.

An example of this form of decay occurs in the **uranium-238** nucleus. The equation which represents what occurs is:



Here the uranium-238 nucleus emits a helium-4 nucleus (the alpha particle) and the parent nucleus becomes thorium-234. Note that the Mass Number of the parent nucleus has been reduced by 4 and the Atomic Number is reduced by 2 which is a characteristic of alpha decay for any nucleus in which it occurs.

Beta Decay

There are three common forms of **beta decay**:

(a) Electron Emission

Certain nuclei which have an excess of neutrons may attempt to reach stability by converting a neutron into a proton with the emission of an electron. The electron is called a **beta-minus particle** - the minus indicating that the particle is negatively charged.

We can represent what occurs as follows:



where a neutron converts into a proton and an electron. Notice that the total electrical charge is the same on both sides of this equation. We say that the **electric charge is conserved**.

We can consider that the electron cannot exist inside the nucleus and therefore is ejected.

Once again there is nothing strange or mysterious about an electron. What is important though

from a radiation safety point of view is the energy with which it is emitted and the chemical damage it can cause when it interacts with living matter.

An example of this type of decay occurs in the **iodine-131** nucleus which decays into xenon-131 with the emission of an electron, that is



The electron is what is called a beta-minus particle. Note that the Mass Number in the above equation remains the same and that the Atomic Number increases by 1 which is characteristic of this type of decay.

You may be wondering how an electron can be produced inside a nucleus given that the simple atomic description we gave in the previous chapter indicated that the nucleus consists of protons and neutrons only. This is one of the limitations of the simple treatment presented so far and can be explained by considering that the three fundamental particles which we call protons, neutrons and electrons are themselves formed of smaller particles called **quarks**. We are not going to consider these in any way here other than to note that some combinations of different types of **quark** produce protons, another combination produces neutrons and a third combination forms electrons. The message here is to appreciate that a simple picture is the best way to start in an introductory text such as this and that the real situation is a lot more complex than what has been described. The same can be said about the treatment of beta-decay given above as we will see in subsequent chapters.

(b) Positron Emission

When the number of protons in a nucleus is too large for the nucleus to be stable it may attempt to reach stability by converting a proton into a neutron with the emission of a positively-charged electron.

That is not a typographical error! An electron with a positive charge also called a **positron** is emitted. The positron is the **beta-plus particle**.

The history here is quite interesting. A brilliant Italian physicist, **Enrico Fermi** developed a theory of beta decay and his theory predicted that positively-charged as well as negatively-charged electrons could be emitted by unstable nuclei. These particles could be called pieces of **anti-matter** and they were subsequently discovered by experiment. They do not exist for very long as they quickly combine with a normal electron and the subsequent reaction called **annihilation** gives rise to the emission of two gamma rays.

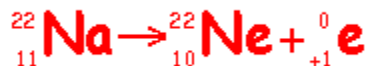
Science fiction writers had a great time following the discovery of anti-matter and speculated along with many scientists that parts of our universe may contain negatively-charged protons forming nuclei which are orbited by positively-charged electrons. But this is taking us too far away from the topic at hand!

The reaction in our unstable nucleus which contains one too many protons can be represented as follows:

Notice, once again, that electric charge is conserved on each side of this equation.



An example of this type of decay occurs in **sodium-22** which decays into **neon-22** with the emission of a positron:



Note that the Mass Number remains the same and that the Atomic Number decreases by 1.

(c) Electron Capture

In this third form of beta decay an inner orbiting electron is attracted into an unstable nucleus where it combines with a proton to form a neutron. The reaction can be represented as:

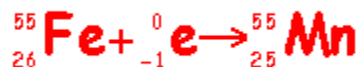


This process is also known as **K-capture** since the electron is often attracted from the K-shell of the atom.

How do we know that a process like this occurs given that no radiation is emitted? In other words the event occurs within the atom itself and no information about it leaves the atom. Or does it? The signature of this type of decay can be obtained from effects in the electron cloud surrounding the nucleus when the vacant site left in the K-shell is filled by an electron from an outer shell. The filling of the vacancy is associated with the **emission of an X-ray** from the electron cloud and it is this X-ray which provides a signature for this type of beta decay.

This form of decay can also be recognised by the emission of gamma-rays from the new nucleus.

An example of this type of radioactive decay occurs in **iron-55** which decays into **manganese-55** following the capture of an electron. The reaction can be represented as follows:



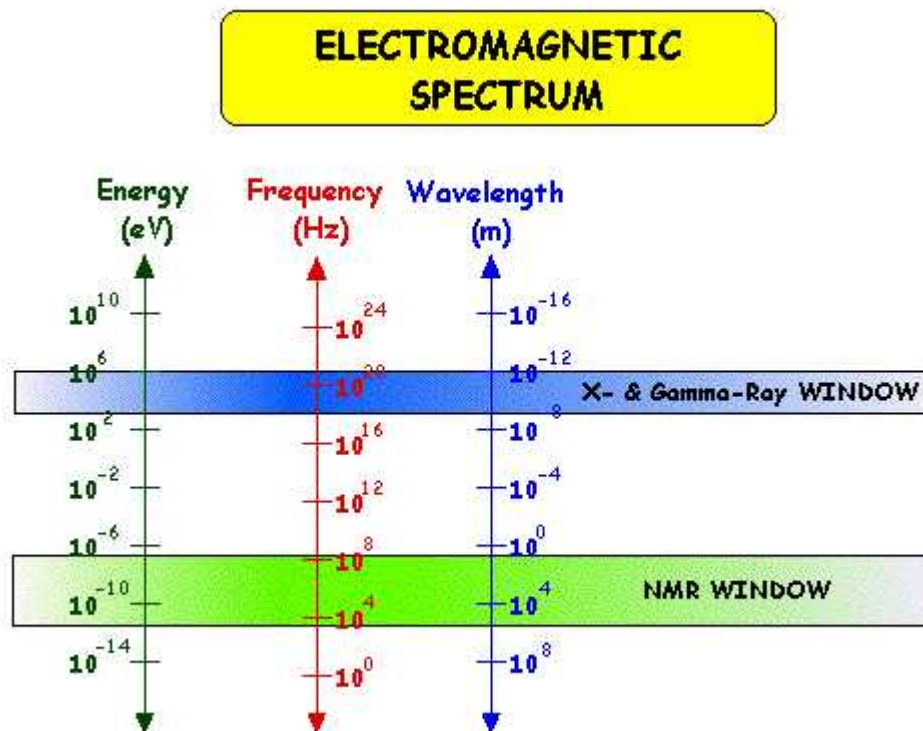
Note that the Mass Number once again is unchanged in this form of decay and that the Atomic Number is decreased by 1.

Gamma Decay

Gamma decay involves the emission of energy from an unstable nucleus in the form of electromagnetic radiation.

You should remember from your high school physics that **electromagnetic radiation** is the biggest physical phenomenon we have so far discovered. The radiation can be characterised in terms of its frequency, its wavelength and its energy. Thinking about it in terms of the energy of the radiation we

have very low energy electromagnetic radiation called **radio waves**, **infra-red radiation** at a slightly higher energy, **visible light** at a higher energy still, then **ultra-violet radiation** and the higher energy forms of this radiation are called **X-rays** and **gamma-rays**. You should also remember that these radiations form what is called the Electromagnetic Spectrum.



Before proceeding it is useful to pause for a moment to consider the difference between X-rays and gamma-rays. These two forms of radiation are high energy electromagnetic rays and are therefore virtually the same. The difference between them is not **what** they consist of but **where** they come from. In general we can say that if the radiation emerges from a nucleus it is called a gamma-ray and if it emerges from outside the nucleus from the electron cloud for example, it is called an X-ray.

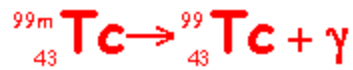
One final point is of relevance before we consider the different forms of gamma-decay and that is what such a high energy ray really is. It has been found in experiments that gamma-rays (and X-rays for that matter!) sometimes manifest themselves as waves and other times as particles. This **wave-particle duality** can be explained using the equivalence of mass and energy at the atomic level. When we describe a gamma ray as a wave it has been found useful to use terms such as frequency and wavelength just like any other wave. In addition when we describe a gamma ray as a particle we use terms such as mass and electric charge. Furthermore the term **electromagnetic photon** is used for these particles. The interesting feature about these photons however is that they have neither mass nor charge!

There are two common forms of gamma decay:

(a) Isomeric Transition

A nucleus in an excited state may reach its ground or unexcited state by the emission of a gamma-ray.

An example of this type of decay is that of **technetium-99m** - which by the way is the most common radioisotope used for diagnostic purposes today in medicine. The reaction can be expressed as:



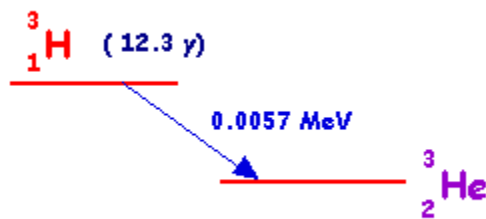
Here a nucleus of technetium-99 is in an excited state, that is it has excess energy. The excited state in this case is called a **metastable state** and the nucleus is therefore called technetium-99m (m for metastable). The excited nucleus loses its excess energy by emitting a gamma-ray to become technetium-99.

(b) Internal Conversion

Here the excess energy of an excited nucleus is given to an atomic electron, e.g. a K-shell electron.

Decay Schemes

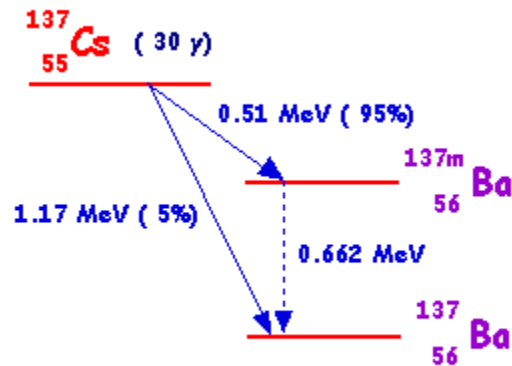
Decay schemes are widely used to give a visual representation of radioactive decay. A scheme for a relatively straight-forward decay is shown below:



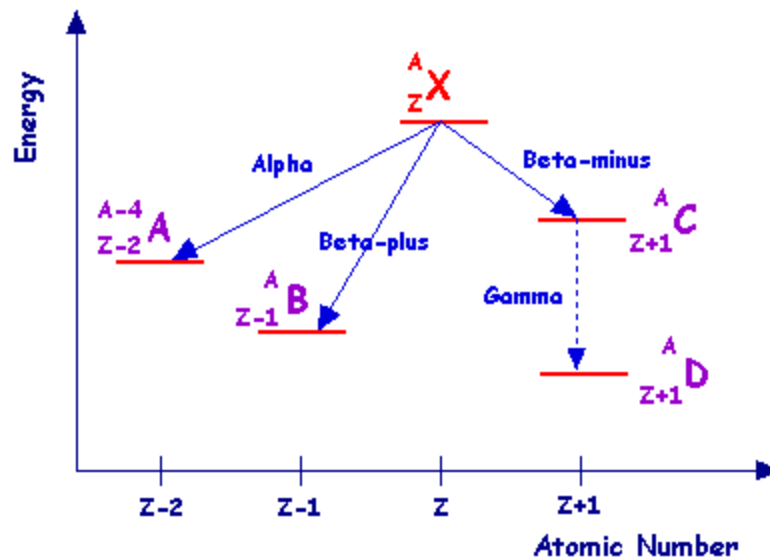
This scheme is for hydrogen-3 which decays to helium-3 with a half-life of 12.3 years through the emission of a beta-minus particle with an energy of 0.0057 MeV.

A scheme for a more complicated decay is that of caesium-137:

This isotope can decay through through two beta-minus processes. In one which occurs in 5% of disintegrations a beta-minus particle is emitted with an energy of 1.17 MeV to produce barium-137. In the second which occurs more frequently (in the remaining 95% of disintegrations) a beta-minus particle of energy 0.51 MeV is emitted to produce barium-137m - in other words a barium-137 nucleus in a metastable state. The barium-137m then decays via isomeric transition with the emission of a gamma-ray of energy 0.662 MeV.



The general method used for decay schemes is illustrated in the diagram below:



The energy is plotted on the vertical axis and atomic number on the horizontal axis - although these axes are rarely displayed in actual schemes. The isotope from which the scheme originates is displayed at the top - X in the case above. This isotope is referred to as the **parent**. The parent loses energy when it decays and hence the products of the decay referred to as **daughters** are plotted at a lower energy level.

The diagram illustrates the situation for common forms of radioactive decay. Alpha-decay is illustrated on the left where the mass number is reduced by 4 and the atomic number is reduced by 2 to produce daughter **A**. To its right the scheme for beta-plus decay is shown to produce daughter **B**. The situation for beta-minus decay followed by gamma-decay is shown on the right side of the diagram where daughters **C** and **D** respectively are produced.

External Links

- [Basics about Radiation](#) - overview of the different types of ionising radiation from the

Radiation Effects Research Foundation - a cooperative Japan-United States Research Organization which conducts research for peaceful purposes.

- **Radiation and Life** - an essay by Eric J Hall, Professor of Radiology, Columbia University with sections on the unstable atom, ionising radiation, background radiation, man-made radiation and radioactive decay - from the Uranium Information Centre website.
- **Radiation and Radioactivity** - a self-paced lesson developed by the University of Michigan's Student Chapter of the Health Physics Society, with sections on radiation, radioactivity, the atom, alpha radiation, beta radiation and gamma radiation.

Multiple Choice Questions

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CHAPTER 02 • RADIOACTIVE DECAY

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CHAPTER 3 THE RADIOACTIVE DECAY LAW

Introduction

We covered radioactive decay from a phenomenological perspective in the last chapter. In this chapter we consider the topic from a more general analytical perspective.

The reason for doing this is so that we can develop a form of thinking which will help us to understand what is going on in a quantitative, mathematical sense. We will be introduced to concepts such as the **Decay Constant** and the **Half Life** as well as units used for the measurement of radioactivity. You will also have a chance to develop your understanding by being brought through three questions on this subject.

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Assumptions

The usual starting point in most forms of analysis in physics is to make some assumptions which simplify the situation. By simplifying the situation we can dispose of irrelevant effects which tend to complicate matters but in doing so we sometimes make the situation so simple that it becomes a bit too abstract and apparently hard to understand.

For this reason we will try here to relate the subject of radioactive decay to a more common situation which we will use as an analogy and hopefully we will be able to overcome the abstract feature of the subject matter. The analogy we will use here is that of making **popcorn**.

So think about putting some oil in a pot, adding the corn, heating the pot on the cooker and watching what happens. You might also like to try this out while considering the situation!

For our radioactive decay situation we first of all consider that we have a sample containing a large number of radioactive nuclei all of the same kind. This is our unpopped corn in the pot for example.

Secondly we assume that all of the radioactive nuclei decay by the same process be it alpha, beta or gamma-decay. In other words our unpopped corn goes pop at some stage during the heating process.

Thirdly take a few moments to ponder on the fact that we can only really consider what is going on from a statistical perspective. If you look at an individual piece of corn, can you figure out when it is going to pop? Not really. You can however figure out that a large number of them will have popped after a period of time. But its rather more difficult to figure out the situation for an individual piece of corn. So instead of dealing with individual entities we consider what happens on a larger scale and this is where **statistics** comes in. We can say that the radioactive decay is a statistical one-shot process, that is when a nucleus has decayed it cannot repeat the process again. In other words when a piece of corn has popped it cannot repeat the process. Simple!

In addition as long as a radioactive nucleus has not decayed the probability for it doing so in the next moment remains the same. In other words if a piece of corn has not popped at a certain time the

chance of it popping in the next second is the same as in the previous second. The bets are even!

Let us not push this popcorn analogy too far though in that we know that we can control the rate of popping by the heat we apply to the pot for example. However as far as our radioactive nuclei are concerned there is nothing we can do to control what is going on. The rate at which nuclei go pop (or **decay**, in other words) cannot be influenced by heating up the sample. Nor by cooling it for that matter or by putting it under greater pressures, by changing the gravitational environment by taking it out into space for instance, or by changing any other aspect of its physical environment. The only thing that determines whether an individual nucleus will decay seems to be the nucleus itself. But on the average we can say that it will decay at some stage.

The Radioactive Decay Law

Let us now use some symbols to reduce the amount of writing we have to do to describe what is going on and to avail ourselves of some mathematical techniques to simplify the situation even further than we have been able to do so far.

Let us say that in the sample of radioactive material there are N nuclei which have not decayed at a certain time, t . So what happens in the next brief period of time? Some nuclei will decay for sure. But how many?

On the basis of our reasoning above we can say that the number which will decay will depend on overall number of nuclei, N , and also on the length of the brief period of time. In other words the more nuclei there are the more will decay and the longer the time period the more nuclei will decay. Let us denote the number which will have decayed as dN and the small time interval as dt .

So we have reasoned that the number of radioactive nuclei which will decay during the time interval from t to $t+dt$ must be proportional to N and to dt . In symbols therefore:

$$-dN \propto N \cdot dt$$

the minus sign indicating that N is decreasing.

Turning the proportionality in this equation into an equality we can write:

$$-dN = \lambda N \cdot dt$$

where the constant of proportionality, λ , is called the **Decay Constant**.

Dividing across by N we can rewrite this equation as:

$$-\frac{dN}{N} = \lambda \cdot dt$$

So this equation describes the situation for any brief time interval, dt . To find out what happens for all periods of time we simply add up what happens in each brief time interval. In other words we **integrate** the above equation. Expressing this more formally we can say that for the period of time from $t = 0$ to any later time t , the number of radioactive nuclei will decrease from N_0 to N_t , so that:

$$-\int_{N_0}^{N_t} \frac{dN}{N} = \lambda \int_0^t dt$$

$$\therefore \ln \left(\frac{N_t}{N_0} \right) = -\lambda t$$

$$\therefore \frac{N_t}{N_0} = \exp(-\lambda t)$$

$$\therefore N_t = N_0 \exp(-\lambda t)$$

This final expression is known as the **Radioactive Decay Law**. It tells us that the number of radioactive nuclei will decrease in an exponential fashion with time with the rate of decrease being controlled by the **Decay Constant**.

Before looking at this expression in further detail let us review the mathematics which we used above. First of all we used integral calculus to figure out what was happening over a period of time by integrating what we knew would occur in a brief interval of time. Secondly we used a calculus relationship that the

$$\int \frac{dx}{x} = \ln x$$

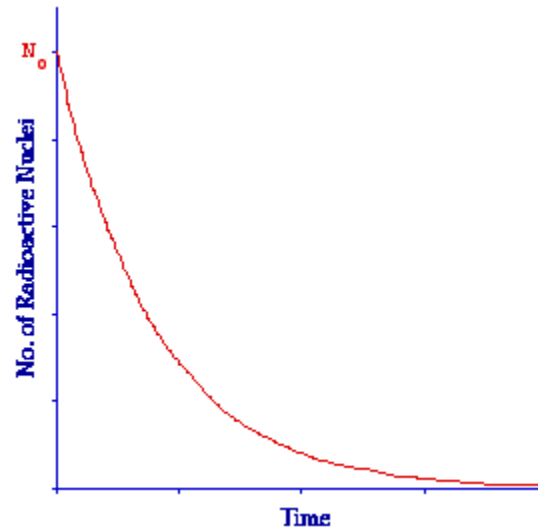
where $\ln x$ represents the **natural logarithm** of x . And thirdly we used the definition of logarithms that when

$$\ln x = y$$

then,

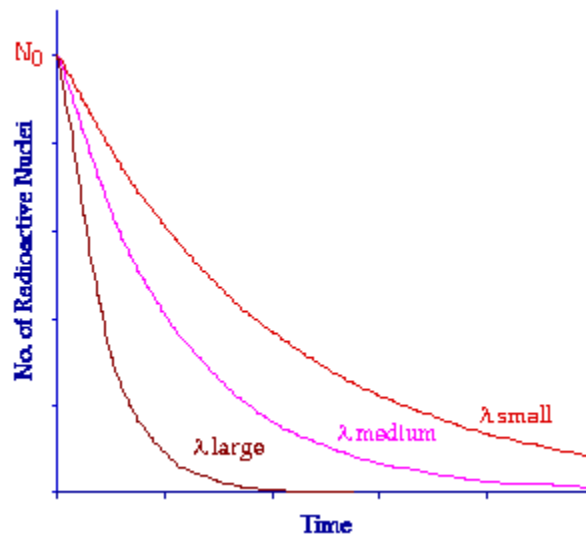
$$x = \exp y$$

Now, to return to the Radioactive Decay Law. The Law tells us that the number of radioactive nuclei will decrease with time in an exponential fashion with the rate of decrease being controlled by the Decay Constant. The Law is shown in graphical form in the figure below:



The graph plots the number of radioactive nuclei at any time, N_t , against time, t . We can see that the number of radioactive nuclei decreases from N_0 that is the number at $t = 0$ in a rapid fashion initially and then more slowly in the classic exponential manner.

The influence of the Decay Constant can be seen in the following figure:



All three curves here are exponential in nature, only the Decay Constant is different. Notice that when the Decay Constant has a low value the curve decreases relatively slowly and when the Decay Constant is large the curve decreases very quickly.

The Decay Constant is characteristic of individual radionuclides. Some like uranium-238 have a small value and the material therefore decays quite slowly over a long period of time. Other nuclei such as technetium-99m have a relatively large Decay Constant and they decay far more quickly.

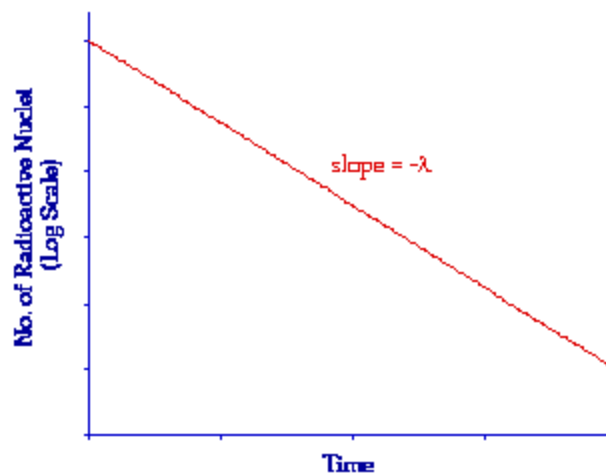
It is also possible to consider the Radioactive Decay Law from another perspective by plotting the logarithm of N_t against time. In other words from our analysis above by plotting the expression:

$$\ln \left(\frac{N_t}{N_0} \right) = -\lambda t$$

in the form

$$\ln N_t = -\lambda t + \ln N_0$$

Notice that this expression is simply an equation of the form $y = mx + c$ where $m = -\lambda$ and $c = \ln N_0$. As a result it is the equation of a straight line of slope $-\lambda$ as shown in the following figure. Such a plot is sometimes useful when we wish to consider a situation without the complication of the direct exponential behaviour.



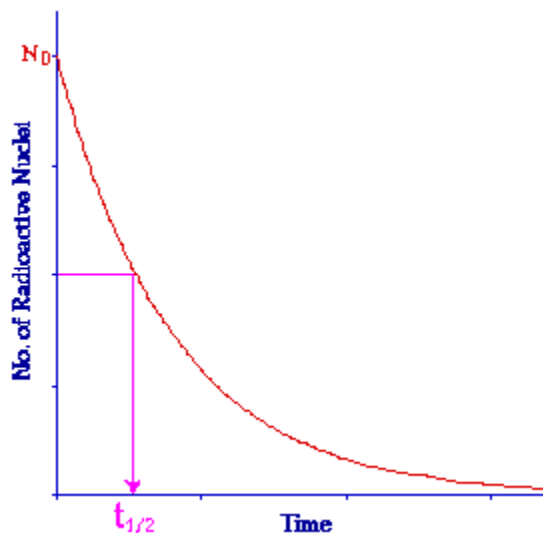
Half-Life

Most of us have not been taught to think instinctively in terms of logarithmic or exponential terms even though many natural phenomena display exponential behaviours. Most of the forms of thinking which we have been taught in school are based on linear changes and as a result it is rather difficult for us to grasp the Radioactive Decay Law intuitively. For this reason an indicator is usually derived from the law which helps us think more clearly about what is going on.

This indicator is called the **Half Life** and it expresses the length of time it takes for the radioactivity of a radioisotope to decrease by a factor of two. From a graphical point of view we can say that when:

$$N_t = \frac{N_0}{2}$$

the time taken is the Half Life:



Note that the **half-life** does not express how long a material will remain radioactive but simply the length of time for its radioactivity to halve. Examples of the half lives of some radioisotopes are given in the following table. Notice that some of these have a relatively short half life. These tend to be the ones used for medical diagnostic purposes because they do not remain radioactive for very long following administration to a patient and hence result in a relatively low radiation dose.

Radioisotope	Half Life (approx.)
$^{81\text{m}}\text{Kr}$	13 seconds
$^{99\text{m}}\text{Tc}$	6 hours
^{131}I	8 days
^{51}Cr	1 month
^{137}Cs	30 years
^{241}Am	462 years
^{226}Ra	1620 years
^{238}U	4.51×10^9 years

But they do present a logistical problem when we wish to use them when there may not be a radioisotope production facility nearby. For example suppose we wish to use $^{99\text{m}}\text{Tc}$ for a patient study and the nearest nuclear facility for making this isotope is 5,000 km away. The production facility could

be in Sydney and the patient could be in Perth for instance. After making the isotope at the nuclear plant it would be decaying with a half life of 6 hours. So we put the material on a truck and drive it to Sydney airport. The isotope would be decaying as the truck sits in Sydney traffic then decaying still more as it waits for a plane to take it to Perth. Then decaying more as it is flown across to Perth and so on. By the time it gets to our patient it will have substantially reduced in radioactivity possibly to the point of being useless for the patient's investigation. And what about the problem if we needed to use ^{81m}Kr instead of ^{99m}Tc for our patient? You will see in the last chapter of this wikibook that logistical challenges such as this have given rise to quite innovative solutions. More about that later!

You can appreciate from the table above that other isotopes have a very long half lives. For example ^{226}Ra has a half life of over 1,500 years. This isotope has been used in the past for therapeutic applications in medicine. Think about the logistical problems here. They obviously do not relate to transporting the material from the point of production to the point of use. But they relate to how the material is kept following its arrival at the point of use. We must have a storage facility so that the material can be kept safely for a long period of time. But for how long? A general rule of thumb for the quantities of radioactivity used in medicine is that the radioactivity will remain significant for about 10 half lives. So we would have to have a safe environment for storage of the ^{226}Ra for about 16,000 years! This storage facility would have to be secure from many unforeseeable events such as earthquakes, bombing etc and be kept in a manner which our children's, children's children can understand. A very serious undertaking indeed!

Relationship between the Decay Constant and the Half Life

On the basis of the above you should be able to appreciate that there is a relationship between the Decay Constant and the Half Life. For example when the Decay Constant is small the Half Life should be long and correspondingly when the Decay Constant is large the Half Life should be short. But what exactly is the nature of this relationship?

We can easily answer this question by using the definition of Half Life and applying it to the Radioactive Decay Law.

Once again the law tells us that at any time, t :

$$N_t = N_0 \exp(-\lambda t)$$

and the definition of Half Life tells us that:

$$N_t = \frac{N_0}{2}$$

when

$$t = t_{\frac{1}{2}}$$

We can therefore re-write the Radioactive Decay Law by substituting for N_t and t as follows:

Therefore

$$\frac{N_0}{2} = N_0 \exp(-\lambda t_{\frac{1}{2}})$$

$$\therefore 2^{-1} = \exp(-\lambda t_{\frac{1}{2}})$$

$$\therefore \ln 2^{-1} = -\lambda t_{\frac{1}{2}}$$

$$\therefore \ln 2 = \lambda t_{\frac{1}{2}}$$

$$\therefore 0.693 = \lambda t_{\frac{1}{2}}$$

$$t_{\frac{1}{2}} = \frac{0.693}{\lambda}$$

and

$$\lambda = \frac{0.693}{t_{\frac{1}{2}}}$$

These last two equations express the relationship between the Decay Constant and the Half Life. They are very useful as you will see when solving numerical questions relating to radioactivity and usually form the first step in solving a numerical problem.

Units of Radioactivity

The SI or metric unit of radioactivity is named after Henri Becquerel, in honour of his discovery of radioactivity, and is called the **becquerel** with the symbol Bq. The **becquerel** is defined as the quantity of radioactive substance that gives rise to a decay rate of 1 decay per second.

In medical diagnostic work 1 Bq is a rather small amount of radioactivity. Indeed it is easy to remember its definition if you think of it as a *buggerall* amount of radioactivity. For this reason the kilobecquerel (kBq) and megabecquerel (MBq) are more frequently used.

The traditional unit of radioactivity is named after Marie Curie and is called the **curie**, with the symbol Ci. The **curie** is defined as the amount of radioactive substance which gives rise to a decay rate

of 3.7×10^{10} decays per second. In other words 37 thousand, million decays per second which as you might appreciate is a substantial amount of radioactivity. For medical diagnostic work the millicurie (mCi) and the microcurie (μCi) are therefore more frequently used.

Why two units? It in essence like all other units of measurement depends on what part of the world you are in. For example the kilometer is widely used in Europe and Australia as a unit of distance and the mile is used in the USA. So if you are reading an American textbook you are likely to find the curie used as the unit of radioactivity, if you are reading an Australian book it will most likely refer to becquerels and both units might be used if you are reading a European book. You will therefore find it necessary to know and understand both units.

Questions

Three questions are given below to help you develop your understanding of the material presented in this chapter. The first one is relatively straight-forward and will exercise your application of the Radioactive Decay Law as well as your understanding of the concept of Half Life. The second question is a lot more challenging and will help you relate the Radioactive Decay Law to the number of radioactive nuclei which are decaying in a sample of radioactive material. The third question will help you understand the approach used in the second question by asking a similar question from a slightly different perspective.

Before you try the questions click **HERE** to access an interactive webpage which will do some simple calculations for you.

Question 1

- (a) The half-life of $^{99\text{m}}\text{Tc}$ is 6 hours. After how much time will 1/16th of the radioisotope remain?
- (b) Verify your answer by another means.

Answer:

(a) Starting with the relationship we established earlier between the Decay Constant and the Half Life we can calculate the Decay Constant as follows:

$$\lambda = \frac{0.693}{t_{\frac{1}{2}}} = \frac{0.693}{6} = 0.1155 \text{ hr}^{-1}$$

Now applying the Radioactive Decay Law,

$$N_t = N_0 \exp(-\lambda t)$$

we can re-write it in the form:

The question tells us that N_0 has reduced to 1/16th of its value, that is:

$$\frac{N_t}{N_0} = \exp(-\lambda t)$$

Therefore

$$\frac{1}{16} = \exp(-0.1155t)$$

which we need to solve for t. One way of doing this is as follows:

$$16^{-1} = \exp(-0.1155t)$$

$$t \therefore \frac{-\ln 16}{0.1155} = \frac{-0.1155t}{24 \text{ hours}}$$

So it will take 24 hours until 1/16th of the radioactivity remains.

(b) A way in which this answer can be verified is by using the definition of Half Life. We are told that the Half Life of ^{99m}Tc is 6 hours. Therefore after six hours half of the radioactivity remains.

Therefore after 12 hours a quarter remains; after 18 hours an eighth remains and after 24 hours one sixteenth remains. And we arrive at the same answer as in part (a). So we must be right!

Note that this second approach is useful if we are dealing with relatively simple situations where the radioactivity is halved, quartered and so on. But supposing the question asked how long would it take for the radioactivity to decrease to a tenth of its initial value. Deduction from the definition of half life is rather more difficult in this case and the mathematical approach used for part (a) above will yield the answer more readily.

Question 2

Find the radioactivity of a 1 g sample of ^{226}Ra given that $t_{1/2}$: 1620 years and Avogadro's Number: 6.023×10^{23} .

Answer:

We can start the answer like we did with Question 1(a) by calculating the Decay Constant from the Half Life using the following equation:

$$\lambda = \frac{0.693}{t_{\frac{1}{2}}} = \frac{0.693}{1620} = 4.28 \times 10^{-4} \text{ year}^{-1}$$

$$\therefore \lambda = 1.36 \times 10^{-11} \text{ s}^{-1}$$

Note that the length of a year used in converting from 'per year' to 'per second' above is 365.25 days to account for leap years. In addition the reason for converting to units of 'per second' is because the unit of radioactivity is expressed as the number of nuclei decaying per second.

Secondly we can calculate that 1 g of ^{226}Ra contains:

$$N = \frac{(\text{Avogadro's No.})(\text{Mass})}{\text{Mass Number}} = \frac{(6.023 \times 10^{23})(1\text{g})}{226} = 2.7 \times 10^{21} \text{ nuclei}$$

Thirdly we need to express the Radioactive Decay Law in terms of the number of nuclei decaying per unit time. We can do this by **differentiating** the equation as follows:

$$N = N_0 \exp(-\lambda t)$$

$$\therefore \frac{dN}{dt} = N_0 \cdot -\lambda \exp(-\lambda t) = -\lambda N_0 \exp(-\lambda t)$$

$$\therefore \frac{dN}{dt} = -\lambda N$$

$$\therefore \left| \frac{dN}{dt} \right| = \lambda N$$

The reason for expressing the result above in absolute terms is to remove the minus sign in that we already know that the number is decreasing.

We can now enter the data we derived above for λ and N :

$$\left| \frac{dN}{dt} \right| = (1.36 \times 10^{-11})(2.7 \times 10^{21})$$

$$\therefore \left| \frac{dN}{dt} \right| = 3.6 \times 10^{10} \text{ decays per second}$$

So the radioactivity of our 1 g sample of radium-226 is approximately 1 Ci.

This is not a surprising answer since the definition of the **curie** was originally conceived as the radioactivity of 1 g of radium-226!

Question 3

What is the minimum mass of ^{99m}Tc that can have a radioactivity of 1 MBq? Assume the half-life is 6 hours and that Avogadro's Number is 6.023×10^{23} .

Answer

Starting again with the relationship between the Decay Constant and the Half Life:

$$\lambda = \frac{0.693}{6} = 0.1155 \text{ hour}^{-1} = 3.21 \times 10^{-5} \text{ s}^{-1}$$

Secondly the question tells us that the radioactivity is 1 MBq. Therefore since $1 \text{ MBq} = 1 \times 10^6$ decays per second,

$$\left| \frac{dN}{dt} \right| = \lambda N = 1 \times 10^6 \text{ dps}$$

Finally the mass of these nuclei can be calculated as follows:

$$\text{Mass of N nuclei} = \frac{(\text{No. of Nuclei})(\text{Mass No.})}{\text{Avogadro's Number}}$$

$$= \frac{(3.116 \times 10^{10})(99)}{6.023 \times 10^{23}} = 5.122 \times 10^{-12} \text{ g}$$

In other words a mass of just over five picograms of ^{99m}Tc can emit one million gamma-rays per second. The result reinforces an important point that you will learn about radiation protection which is that you should treat radioactive materials just like you would handle pathogenic bacteria!

Multiple Choice Questions

Click [HERE](#) to access an online MCQ covering the material in this chapter.

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CHAPTER 4 UNITS OF RADIATION MEASUREMENT

Introduction

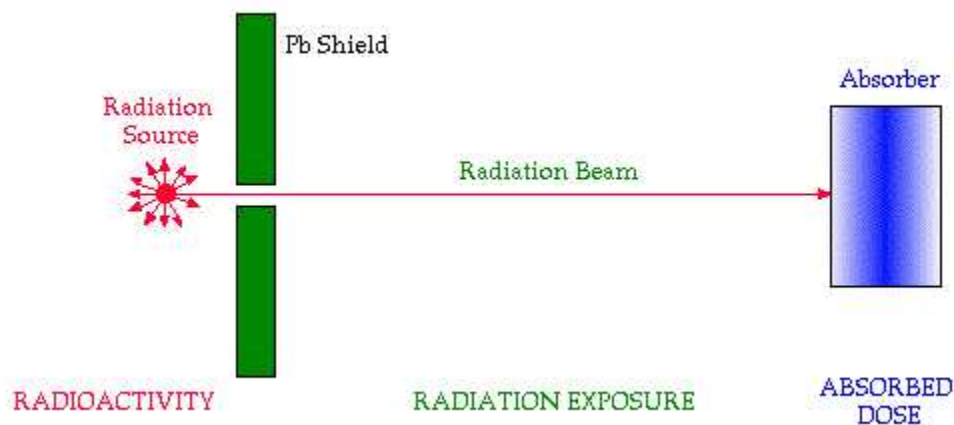
After that rather long and detailed chapter we have just finished we will now proceed at a more leisurely pace for a short treatment of some of the more common units of measurement used in this field.

Before we do so however it is useful to consider the typical radiation environment. By doing so we will gain an appreciation of the various quantities that can be measured before considering the units which are used to express such measurements. So, we will first of all consider a typical radiation situation and then go on to consider the various units of measurement.

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A Typical Radiation Situation

A typical radiation set-up is shown in the figure below. Firstly there is a **source** of radiation, secondly a radiation **beam** and thirdly some **material** which absorbs the radiation. So the quantities which can be measured are associated with the source, the radiation beam and the absorber.



This type of environment could be one where the radiation from the source is used to irradiate a patient (that is the absorber) for diagnostic purposes where we would place a device behind the patient for producing an image or for therapeutic purposes where the radiation is intended to cause damage to a specific region of a patient. It is also a situation where we as an absorber may be working with a source of radiation.

The Radiation Source

When the radiation source is a radioactive one the quantity that is typically measured is the

radioactivity of the source. We saw in the previous chapter that the units used to express radioactivity are the **becquerel** (SI unit) and the **curie** (traditional unit).

The Radiation Beam

The characteristic of a radiation beam that is typically measured is called the **Radiation Exposure**. This quantity expresses how much ionisation the beam causes in the air through which it travels.

We will see in the following chapter that one of the major things that happens when radiation encounters matter is that **ions** are formed - air being the form of matter it encounters in this case. So the radiation exposure produced by a radiation beam is expressed in terms of the amount of ionisation which occurs in air.

A straight-forward way of measuring such ionisation is to determine the amount of electric charge which is produced. You will remember from your high school physics that the SI unit of electric charge is the **coulomb**.

The SI unit of radiation exposure is the **coulomb per kilogram** - and is given the symbol C kg^{-1} . It is defined as the quantity of X- or gamma-rays such that the associated electrons emitted per kilogram of air at standard temperature and pressure (STP) produce ions carrying 1 coulomb of electric charge.

The traditional unit of radiation exposure is the **roentgen**, named in honour of **Wilhelm Roentgen** (who discovered X-rays) and is given the symbol R. The **roentgen** is defined as the quantity of X- or gamma-rays such that the associated electrons emitted per kilogram of air at STP produce ions carrying 2.58×10^{-4} coulombs of electric charge.

So 1 R is a small exposure relative to 1 C kg^{-1} - in fact it is 3,876 times smaller.

Note that this unit is confined to radiation beams consisting of X-rays or gamma-rays.

Often it is not simply the exposure that is of interest but the exposure rate, that is the exposure per unit time. The units which tend to be used in this case are the $\text{C kg}^{-1} \text{ s}^{-1}$ and the R hr^{-1} .

The Absorber

Energy is deposited in the absorber when radiation interacts with it. It is usually quite a small amount of energy but energy nonetheless. The quantity that is measured is called the **Absorbed Dose** and it is of relevance to all types of radiation be they X- or gamma-rays, alpha- or beta-particles.

The SI unit of absorbed dose is called the **gray**, named after a famous radiobiologist, **LH Gray**, and is given the symbol Gy. The **gray** is defined as the absorption of 1 joule of radiation energy per kilogram of material. So when 1 joule of radiation energy is absorbed by a kilogram of the absorber material we say that the absorbed dose is 1 Gy.

The traditional unit of absorbed dose is called the **rad**, which supposedly stands for **Radiation Absorbed Dose**. It is defined as the absorption of 10^{-2} joules of radiation energy per kilogram of

material.

As you can figure out 1 Gy is equal to 100 rad.

There are other quantities derived from the gray and the rad which express the biological effects of such absorbed radiation energy when the absorber is living matter - human tissue for example. These quantities include the **Equivalent Dose** and the **Effective Dose**.

Let us pause here for a bit to ponder on the use of the term *dose*. It usually has a medical connotation in that we can say that someone had a dose of the 'flu, or that the doctor prescribed a certain dose of a drug. What has it to do with the deposition of energy by a beam of radiation in an absorber? It could have something to do with the initial applications of radiation in the early part of the 20th century when it was used to treat numerous diseases. As a result we can speculate that the term has stayed in the vernacular of the field. It would be much easier to use a term like absorbed radiation energy since we are talking about the deposition of energy in an absorber. But this might make the subject just a little too simple!

Specific Gamma Ray Constant

A final quantity is worth mentioning with regard to radiation units. This is the Specific Gamma-Ray Constant for a radioisotope. This quantity is an amalgam of the quantities we have already covered and expresses the exposure rate produced by the gamma-rays emitted from a radioisotope.

It is quite a useful quantity from a practical viewpoint when we are dealing with a radioactive source which emits gamma-rays. Supposing you are using a gamma-emitting radioactive source (for example ^{99m}Tc or ¹³⁷Cs) and you will be standing at a certain distance from this source while you are working. You most likely will be interested in the exposure rate produced by the source from a radiation safety point of view. This is where the Specific Gamma-Ray Constant comes in.

It is defined as the exposure rate per unit activity at a certain distance from a source. The SI unit is therefore the

$$C \text{ kg}^{-1} \text{ s}^{-1} \text{ Bq}^{-1} \text{ at } 1 \text{ m} ,$$

and the traditional unit is the

$$R \text{ hr}^{-1} \text{ mCi}^{-1} \text{ at } 1 \text{ cm} .$$

These units of measurement are quite cumbersome and a bit of a mouthful. It might have been better if they were named after some famous scientist so that we could call the SI unit 1 **smith** and the traditional unit 1 **jones** for example. But again things are not that simple!

The Inverse Square Law

Before we finish this chapter we are going to consider what happens as we move our absorber away from the radiation source. In other words we are going to think about the influence of distance on

the intensity of the radiation beam. You will find that a useful result emerges from this that has a very important impact on radiation safety.

The radiation produced in a radioactive source is emitted in all directions. We can consider that spheres of equal radiation intensity exist around the source with the number of photons/particles spreading out as we move away from the source.

Consider an area on the surface of one of these spheres and assume that there are a certain number of photons/particles passing through it. If we now consider a sphere at a greater distance from the source the same number of photons/particles will now be spread out over a bigger area. Following this line of thought it is easy to appreciate that the radiation intensity will decrease with the square of the distance from the source.

This effect is known as the **Inverse Square Law**. As a result if we double the distance from a source, we reduce the intensity by a factor of two squared, that is 4. If we treble the distance the intensity is reduced by a factor of 9, that is three squared, and so on.

This is a very useful piece of information if you are working with a source of radiation and are interested in minimising the dose of radiation you will receive.

External Links

- **Radiation and Risk** - covers the effect of radiation, how risks are determined, comparison of radiation with other risks and radiation doses.
- **Radiation Effects Overview** - results of studies of victims of nuclear bombs including early effects on survivors, effects on the in utero exposed, and late effects on the survivors - from the Radiation Effects Research Foundation, a cooperative Japan-United States Research Organization.
- **The Radiation and Health Physics Home Page** - all you ever wanted to know about radiation but were afraid to ask....with hundreds of WWW links - from the Student Chapter of the Health Physics Society, University of Michigan containing sections on general information, regulatory information, professional organizations and societies, radiation specialties, health physics research and education.
- **What You Need to Know about Radiation** - to protect yourself to protect your family to make reasonable social and political choices - covers sources of radiation and radiation protection - by Lauriston S. Taylor.

Multiple Choice Questions

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CHAPTER 04 • UNITS OF RADIATION MEASUREMENT

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CHAPTER 5 INTERACTION OF RADIATION WITH MATTER

Introduction

We have focussed in previous chapters on the source of radiation and the types of radiation. We are now in a position to consider what happens when this radiation interacts with matter. Our main reason for doing this is to find out what happens to the radiation as it passes through matter and also to set ourselves up for considering how to detect radiation. Since all radiation detectors are made from some form of matter it is useful to first of all know how radiation interacts so that we can exploit the effects in the design of such detectors.

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Before we do this let us first remind ourselves of the physical characteristics of the major types of radiation. We have covered this material in some detail in the first chapter and it are summarised below for convenience:

Radiation	Mass	Electric Charge	Velocity
Alpha Particles	relatively heavy	double positive	relatively slow
Beta Particles	about 8,000 times lighter	negative	less than the velocity of light
Gamma Rays	None	None	3×10^8 m/s in free space

We will now consider the passage of each type of radiation through matter with most attention given to gamma-rays because they are the most common type used in nuclear medicine. One of the main effects that you will notice irrespective of the type of radiation is that ions are produced when radiation interacts with matter. It is for this reason that it is called **ionizing radiation**.

Before we start though you might find an analogy useful to help you with your thinking. This analogy works on the basis of thinking about matter as an enormous mass of atoms (that is nuclei with orbiting electrons) and that the radiation is a particle/photon passing through this type of environment. So the analogy to think about is a spaceship passing through a meteor storm like you might see in a science-fiction movie where the spaceship represents the radiation and the meteors represent the atoms of the material through which the radiation is passing. One added feature to bring on board however is that our spaceship sometimes has an electric charge depending on the type of radiation.

Alpha Particles

We can see from the table above that **alpha-particles** have a double positive charge and we can therefore easily appreciate that they will exert considerable electrostatic attraction on the outer orbital electrons of atoms near which they pass. The result is that some electrons will be attracted away from

their parent atoms and that ions will be produced. In other words ionizations occur.

We can also appreciate from the table that alpha-particles are quite massive relative to the other types of radiation and also to the electrons of atoms of the material through which they are passing. As a result they travel in straight lines through matter except for rare direct collisions with nuclei of atoms along their path.

A third feature of relevance here is the energy with which they are emitted. This energy in the case of alpha-particles is always distinct. For example ^{221}Ra emits an alpha-particle with an energy of 6.71 MeV. Every alpha-particle emitted from this radionuclide has this energy. Another example is ^{230}U which emits three alpha-particles with energies of 5.66, 5.82, 5.89 MeV.

Finally it is useful to note that alpha-particles are very damaging biologically and this is one reason why they are not used for in-vivo diagnostic studies. We will therefore not be considering them in any great detail in this wikibook.

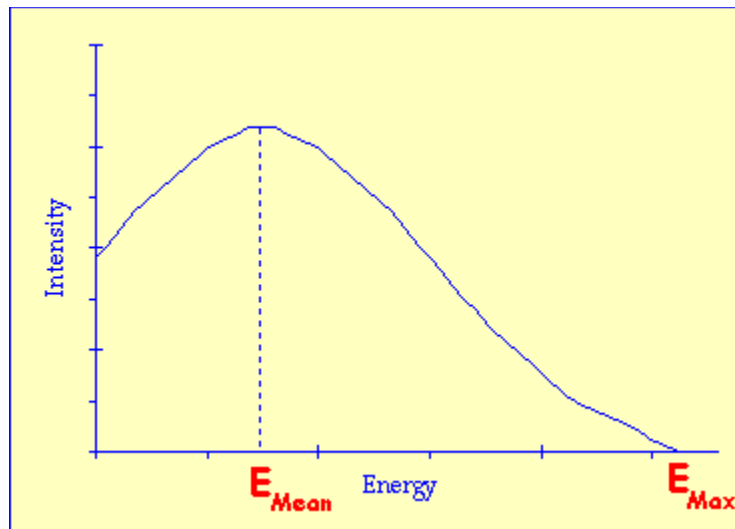
Beta Particles

We can see from the table that **beta-particles** have a negative electric charge. Notice that positrons are not considered here since as we noted in [chapter 2](#) these particles do not last for very long in matter before they are annihilated. Beta-minus particles last considerably longer and are therefore the focus of our attention here.

Because of their negative charge they are attracted by nuclei and repelled by electron clouds as they pass through matter. The result once again without going into great detail is ionization.

The path of beta-particles in matter is often described as being tortuous, since they tend to ricochet from atom to atom.

A final and important point to note is that the energy of beta-particles is never found to be distinct in contrast to the alpha-particles above. The energies of the beta-particles from a radioactive source forms a spectrum up to a maximum energy - see figure below. Notice from the figure that a range of energies is present and features such as the mean energy, E_{mean} , or the maximum energy, E_{max} , are quoted.



The question we will consider here is: why should a spectrum of energies be seen? Surely if a beta-particle is produced inside a nucleus when a neutron is converted into a proton, a single distinct energy should result. The answer lies in the fact that two particles are actually produced in beta-decay. We did not cover this in our treatment in chapter 2 for fear of complicating things too much at that stage of this wikibook. But we will cover it here briefly for the sake of completeness.

The second particle produced in beta-decay is called a **neutrino** and was named by **Enrico Fermi**. It is quite a mysterious particle possessing neither mass nor charge and we are still researching its properties today. The difficulty with them is that they are very hard to detect and this has greatly limited our knowledge about them so far.

The beta-particle energy spectrum can be explained by considering that the energy produced when a neutron is converted to a proton is shared between the beta-particle and the neutrino. Sometimes all the energy is given to the beta-particle and it receives the maximum energy, E_{\max} . But more often the energy is shared between them so that for example the beta-particle has the mean energy, E_{mean} and the neutrino has the remainder of the energy.

Finally it is useful to note that beta-particles are quite damaging biologically and this is one reason why they are not used for in-vivo diagnostic studies. We will therefore not consider them in any great detail in this wikibook.

Gamma Rays

Since we have been talking about energies above let us first note that the energies of **gamma-rays** emitted from a radioactive source are always distinct. For example $^{99\text{m}}\text{Tc}$ emits gamma-rays which all have an energy of 140 keV and ^{51}Cr emits gamma-rays which have an energy of 320 keV.

Gamma-rays have many modes of interaction with matter. Some of those which are not important to nuclear medicine imaging are:

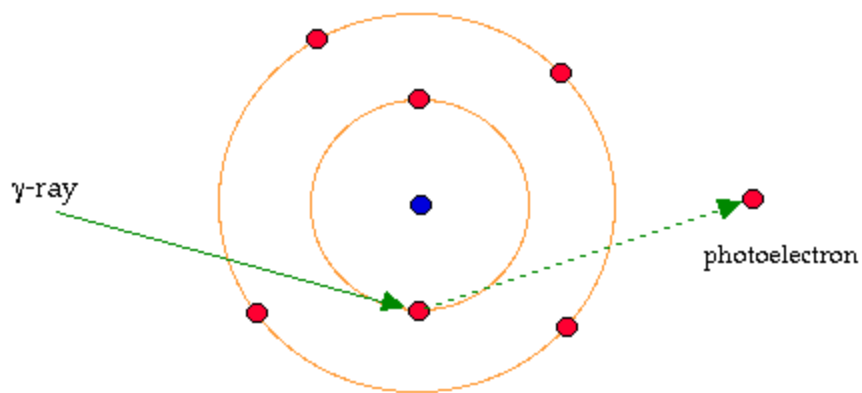
- **Mössbauer Effect**
- **Coherent Scattering**
- **Pair Production**

and will not be described here.

Those which are very important to nuclear medicine imaging, are the **Photoelectric Effect** and the **Compton Effect**. We will consider each of these in turn below. Note that the effects described here are also of relevance to the interaction of X-rays with matter since as we have noted before X-rays and gamma-rays are essentially the same entities. So the treatment below is also of relevance to radiography.

Photoelectric Effect

When a gamma-ray collides with an orbital electron of an atom of the material through which it is passing it can transfer all its energy to the electron and cease to exist - see figure below. On the basis of the **Principle of Conservation of Energy** we can deduce that the electron will leave the atom with a kinetic energy equal to the energy of the gamma-ray less that of the orbital binding energy. This electron is called a **photoelectron**.



Click [HERE](#) for an animated version of this graphic.

Note that an ion results when the photoelectron leaves the atom. Also note that the gamma-ray energy is totally absorbed in the process.

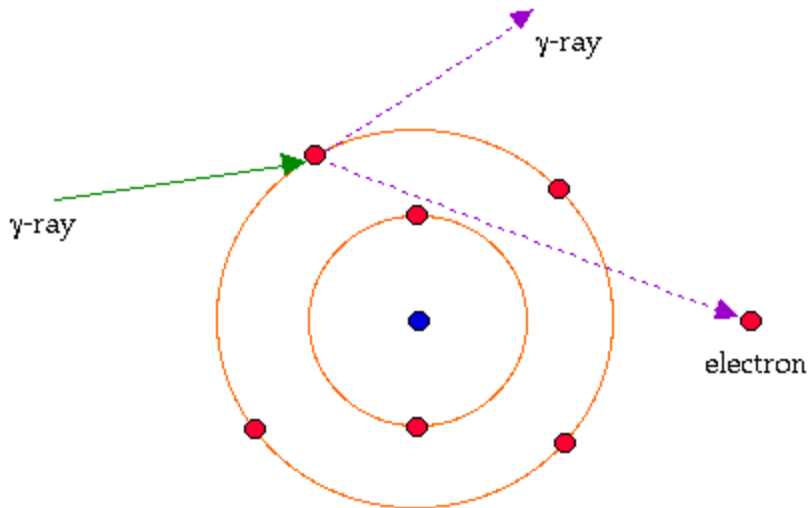
Two subsequent points should also be noted. Firstly the photoelectron can cause ionisations along its track in a similar manner to a beta-particle. Secondly **X-ray emission** can occur when the vacancy left by the photoelectron is filled by an electron from an outer shell of the atom. Remember that we came across this type of feature before when we dealt with Electron Capture in **Chapter 2**.

Compton Effect

This type of effect is somewhat akin to a cue ball hitting a coloured ball on a pool table. Here a gamma-ray transfers only part of its energy to a valance electron which is essentially free - see figure below. Notice that the electron leaves the atom and may act like a beta-particle and that the gamma-ray deflects off in a different direction to that with which it approached the atom. This deflected or

scattered gamma-ray can undergo further Compton Effects within the material.

Note that this effect is sometimes called **Compton Scattering**.



Click [HERE](#) for an animated version of this graphic.

Attenuation

The two effects we have just described give rise to both absorption and scattering of the radiation beam. The overall effect is referred to as attenuation of gamma-rays. We will investigate this feature from an analytical perspective in the following chapter.

Multiple Choice Questions

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CHAPTER 6 ATTENUATION OF GAMMA-RAYS

Introduction

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We covered the interaction of gamma-rays with matter from a descriptive viewpoint in the previous chapter and we saw that the Compton and Photoelectric Effects were the major mechanisms. We will consider the subject again here but this time from an analytical perspective. This will allow us to develop a more general understanding of the phenomenon.

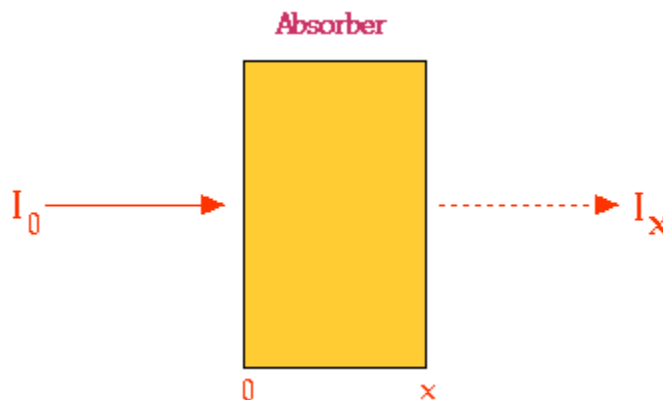
Note that the treatment here also refers to the attenuation of X-rays since, as we noted before gamma-rays and X-rays are essentially the same physical entities.

Our treatment begins with a description of a simple radiation experiment which can be performed easily in the laboratory and which many of the early pioneers in this field did. We will then build on the information obtained from such an experiment to develop a simple equation and some simple concepts which will allow us generalise the situation to any attenuation situation.

Attenuation Experiment

The experiment is quite simple. It involves firing a narrow beam of gamma-rays at a material and measuring how much of the radiation gets through. We can vary the energy of the gamma-rays we use and the type of absorbing material as well as its thickness and density.

The experimental set-up is illustrated in the figure below. We refer to the intensity of the radiation which strikes the absorber as the **incident intensity**, I_0 , and the intensity of the radiation which gets through the absorber as the **transmitted intensity**, I_x . Notice also that the thickness of the absorber is denoted by x .



From what we covered in the previous chapter we can appreciate that some of the gamma-rays will be subjected to interactions such as the Photoelectric Effect and the Compton Effect as they pass through the absorber. The transmitted gamma-rays will in the main be those which pass through

without any interactions at all.

We can therefore expect to find that the transmitted intensity will be less than the incident intensity, that is

$$I_x < I_0$$

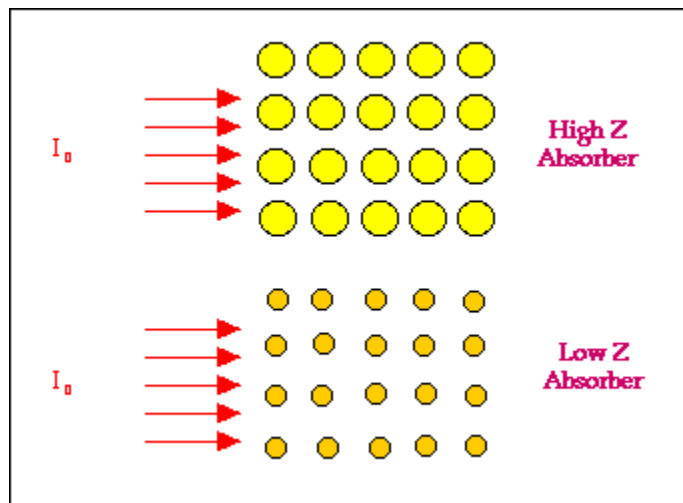
But by how much you might ask. Before we consider this let us denote the difference between I_x and I_0 as ΔI , that is

$$\Delta I = I_0 - I_x$$

Effect of Atomic Number

Let us start exploring the magnitude of ΔI by placing different absorbers in turn in the radiation beam. What we would find is that the magnitude of ΔI is highly dependent on the atomic number of the absorbing material. For example we would find that ΔI would be quite low in the case of an absorber made from **carbon** ($Z=6$) and very large in the case of **lead** ($Z=82$).

We can gain an appreciation of why this is so from the following figure:



The figure illustrates a high atomic number absorber by the large circles which represent individual atoms and a low atomic number material by smaller circles. The incident radiation beam is represented by the arrows entering each absorber from the left. Notice that the atoms of the high atomic number absorber present larger targets for the radiation to strike and hence the chances for interactions via the Photoelectric and Compton Effects is relatively high. The attenuation should therefore be relatively large.

In the case of the low atomic number absorber however the individual atoms are smaller and hence the chances of interactions are reduced. In other words the radiation has a greater probability of being transmitted through the absorber and the attenuation is consequently lower than in the high atomic number case.

With respect to our spaceship analogy used in the previous chapter the atomic number can be thought of as the size of individual meteors in the meteor cloud.

If we were to precisely control our experimental set-up and carefully analyse our results we would find that:

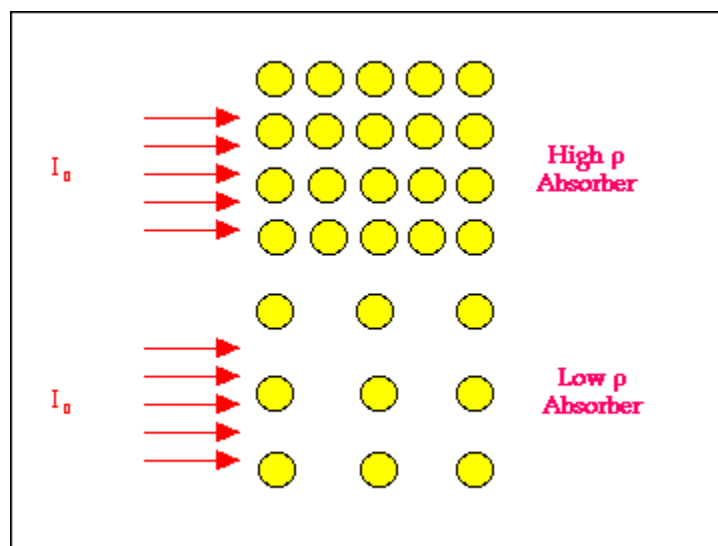
$$\Delta I \propto Z^3$$

Therefore if we were to double the atomic number of our absorber we would increase the attenuation by a factor of two cubed, that is 8, if we were to triple the atomic number we would increase the attenuation by a factor of 27, that is three cubed, and so on.

It is for this reason that high atomic number materials (e.g. Pb) are used for radiation protection.

Effect of Density

A second approach to exploring the magnitude of ΔI is to see what happens when we change the density of the absorber. We can see from the following figure that a low density absorber will give rise to less attenuation than a high density absorber since the chances of an interaction between the radiation and the atoms of the absorber are relatively lower.



So in our analogy of the spaceship entering a meteor cloud think of meteor clouds of different density and the chances of the spaceship colliding with a meteor.

Effect of Thickness

A third factor which we could vary is the thickness of the absorber. As you should be able to predict at this stage the thicker the absorber the greater the attenuation.

Effect of Gamma-Ray Energy

Finally in our experiment we could vary the energy of the gamma-ray beam. We would find without going into it in any great detail that the greater the energy of the gamma-rays the less the attenuation. You might like to think of it in terms of the energy with which the spaceship approaches the meteor cloud and the likelihood of a slow spaceship getting through as opposed to a spaceship travelling with a higher energy.

Mathematical Model

We will consider a mathematical here which will help us to express our experimental observations in more general terms. You will find that the mathematical approach adopted and the result obtained is quite similar to what we encountered earlier with **Radioactive Decay**. So you will not have to plod your way through any new maths below, just a different application of the same form of mathematical analysis!

Let us start quite simply and assume that we vary only the thickness of the absorber. In other words we use an absorber of the same material (i.e. same atomic number) and the same density and use gamma-rays of the same energy for the experiment. Only the thickness of the absorber is changed.

From our reasoning above it is easy to appreciate that the magnitude of ΔI should be dependent on the radiation intensity as well as the thickness of the absorber, that is for an infinitesimally small change in absorber thickness:

$$-dI \propto I \cdot dx$$

the minus sign indicating that the intensity is reduced by the absorber.

Turning the proportionality in this equation into an equality, we can write:

$$-dI = \mu I \cdot dx$$

where the constant of proportionality, μ , is called the **Linear Attenuation Coefficient**.

Dividing across by I we can rewrite this equation as:

$$-\frac{dI}{I} = \mu \cdot dx$$

So this equation describes the situation for any tiny change in absorber thickness, dx . To find out what happens for the complete thickness of an absorber we simply add up what happens in each small

thickness. In other words we integrate the above equation. Expressing this more formally we can say that for thicknesses from $x = 0$ to any other thickness x , the radiation intensity will decrease from I_0 to I_x , so that:

$$-\int_{I_0}^{I_x} \frac{dI}{I} = \mu \int_0^x dx$$

$$\therefore \ln \left(\frac{I_x}{I_0} \right) = -\mu x$$

$$\therefore \frac{I_x}{I_0} = \exp(-\mu x)$$

$$\therefore I_x = I_0 \exp(-\mu x)$$

This final expression tells us that the radiation intensity will decrease in an exponential fashion with the thickness of the absorber with the rate of decrease being controlled by the Linear Attenuation Coefficient. The expression is shown in graphical form below. The graph plots the intensity against thickness, x . We can see that the intensity decreases from I_0 , that is the number at $x = 0$, in a rapid fashion initially and then more slowly in the classic exponential manner.

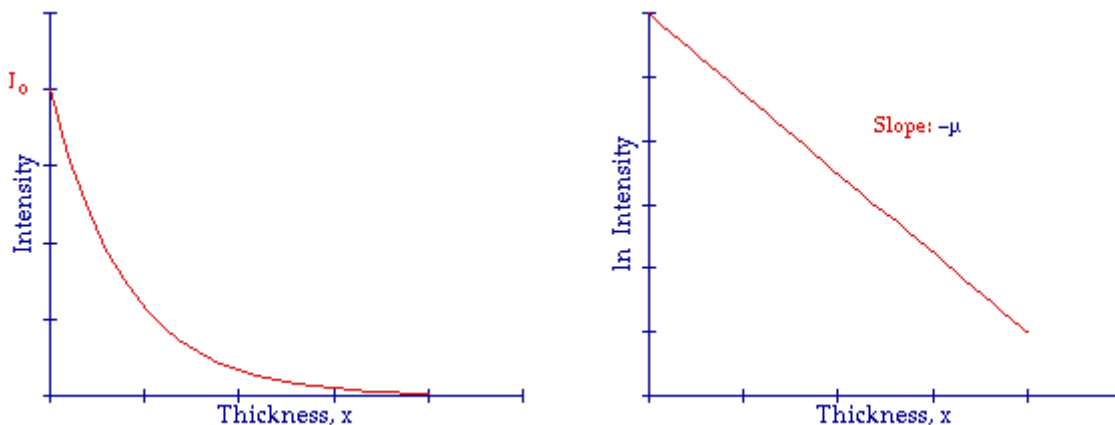


Table 1: Graphical representation of the dependence of radiation intensity on the thickness of absorber: Intensity versus thickness on the left and the natural logarithm of the intensity versus thickness on the right.

The influence of the Linear Attenuation Coefficient can be seen in the next figure. All three curves here are exponential in nature, only the Linear Attenuation Coefficient is different. Notice that when the Linear Attenuation Coefficient has a low value the curve decreases relatively slowly and when the

Linear Attenuation Coefficient is large the curve decreases very quickly.

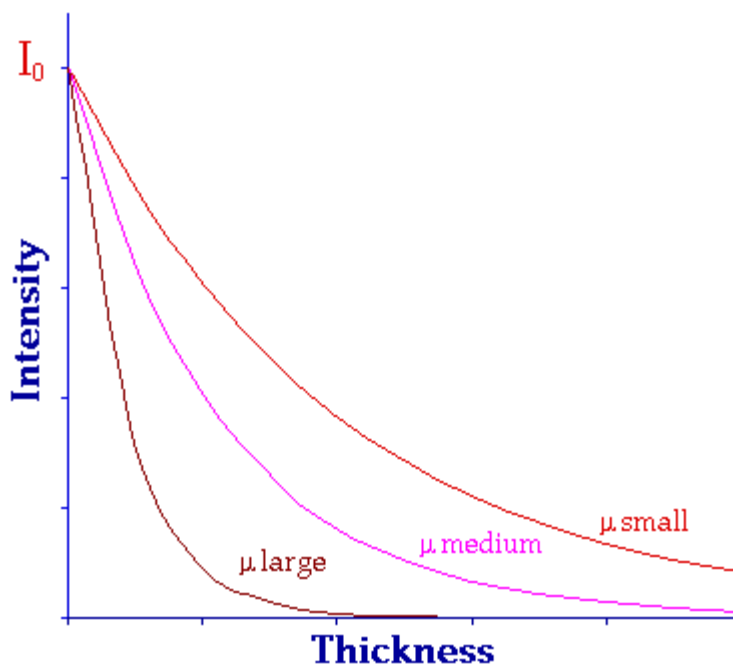


Illustration 1: Exponential attenuation expressed using a small, medium and large value of the Linear Attenuation Coefficient, μ .

The Linear Attenuation Coefficient is characteristic of individual absorbing materials. Some like carbon have a small value and are easily penetrated by gamma-rays. Other materials such as lead have a relatively large Linear Attenuation Coefficient and are relatively good absorbers of radiation:

Absorber	100 keV	200 keV	500 keV
Air	0.000195	0.000159	0.000112
Water	0.167	0.136	0.097
Carbon	0.335	0.274	0.196
Aluminium	0.435	0.324	0.227
Iron	2.72	1.09	0.655
Copper	3.8	1.309	0.73
Lead	59.7	10.15	1.64

Table 2: Linear Attenuation Coefficients (in cm^{-1}) for a range of materials at gamma-ray energies of 100, 200 and 500 keV.

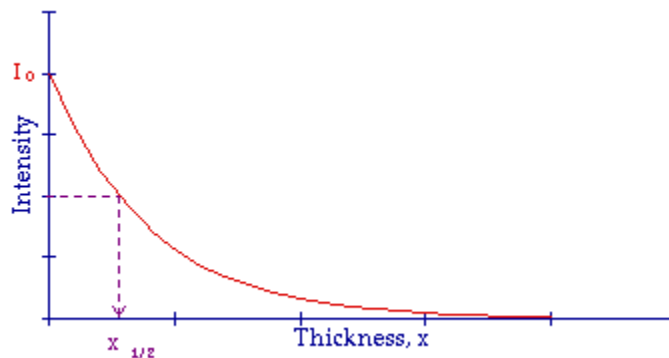
The materials listed in the table above are air, water and a range of elements from carbon ($Z=6$) through to lead ($Z=82$) and their Linear Attenuation Coefficients are given for three gamma-ray energies. The first point to note is that the Linear Attenuation Coefficient increases as the atomic number of the absorber increases. For example it increases from a very small value of 0.000195 cm^{-1} for air at 100 keV to almost 60 cm^{-1} for lead. The second point to note is that the Linear Attenuation Coefficient for all materials decreases with the energy of the gamma-rays. For example the value for copper decreases from about 3.8 cm^{-1} at 100 keV to 0.73 cm^{-1} at 500 keV. The third point to note is that the trends in the table are consistent with the analysis presented earlier. Finally it is important to appreciate that our analysis above is only strictly true when we are dealing with narrow radiation beams. Other factors need to be taken into account when broad radiation beams are involved.

Half Value Layer

As with using the Half Life to describe the **Radioactive Decay Law** an indicator is usually derived from the exponential attenuation equation above which helps us think more clearly about what is going on. This indicator is called the **Half Value Layer** and it expresses the thickness of absorbing material which is needed to reduce the incident radiation intensity by a factor of two. From a graphical point of view we can say that when:

$$I_x = \frac{I_0}{2}$$

the thickness of absorber is the Half Value Layer:



The Half Value Layer for a range of absorbers is listed in the following table for three gamma-ray energies:

Absorber	100 keV	200 keV	500 keV
Air	3,555	4,359	6,189
Water	4.15	5.1	7.15
Carbon	2.07	2.53	3.54

Aluminium	1.59	2.14	3.05
Iron	0.26	0.64	1.06
Copper	0.18	0.53	0.95
Lead	0.012	0.068	0.42

Half Value Layers (in cm) for a range of materials at gamma-ray energies of 100, 200 and 500 keV.

The first point to note is that the Half Value Layer decreases as the atomic number increases. For example the value for air at 100 keV is about 35 meters and it decreases to just 0.12 mm for lead at this energy. In other words 35 m of air is needed to reduce the intensity of a 100 keV gamma-ray beam by a factor of two whereas just 0.12 mm of lead can do the same thing. The second thing to note is that the Half Value Layer increases with increasing gamma-ray energy. For example from 0.18 cm for copper at 100 keV to about 1 cm at 500 keV. Thirdly note that relative to the data in the previous table there is a reciprocal relationship between the Half Value Layer and the Linear Attenuation Coefficient, which we will now investigate.

Relationship between the Linear Attenuation Coefficient and the Half Value Layer

As was the case with the Radioactive Decay Law, where we explored the relationship between the Half Life and the Decay Constant, a relationship can be derived between the Half Value Layer and the Linear Attenuation Coefficient. We can do this by using the definition of the Half Value Layer:

$$I_x = \frac{I_0}{2}$$

when

$$x = x_{\frac{1}{2}}$$

and inserting it in the exponential attenuation equation, that is:

$$I_x = I_0 \exp(-\mu x)$$

to give

Therefore

$$\frac{I_1}{I_2} = \exp(-\mu x_{\frac{1}{2}})$$

$$\therefore 2^{-1} = \exp(-\mu x_{\frac{1}{2}})$$

$$\therefore \ln 2^{-1} = -\mu x_{\frac{1}{2}}$$

$$\therefore \ln 2 = \mu x_{\frac{1}{2}}$$

$$\therefore 0.693 = \mu x_{\frac{1}{2}}$$

$$\mu = \frac{0.693}{x_{\frac{1}{2}}}$$

and

$$x_{\frac{1}{2}} = \frac{0.693}{\mu}$$

These last two equations express the relationship between the Linear Attenuation Coefficient and the Half Value Layer. They are very useful as you will see when solving numerical questions relating to attenuation and frequently form the first step in solving a numerical problem.

Mass Attenuation Coefficient

We implied above that the Linear Attenuation Coefficient was useful when we were considering an absorbing material of the same density but of different thicknesses. A related coefficient can be of value when we wish to include the density, ρ , of the absorber in our analysis. This is the **Mass Attenuation Coefficient** which is defined as the:

$$\frac{\text{Linear Attenuation Coefficient}}{\text{Density}} = \frac{\mu}{\rho}$$

The measurement unit used for the Linear Attenuation Coefficient in the table above is cm^{-1} , and a

common unit of density is the g cm^{-3} . You might like to derive for yourself on this basis that the $\text{cm}^2 \text{g}^{-1}$ is the equivalent unit of the Mass Attenuation Coefficient.

Questions

Two questions are given below to help you develop your understanding of the material presented in this chapter. The first one is relatively straight-forward and will exercise your application of the exponential attenuation equation. The second question is a lot more challenging and will help you relate exponential attenuation to radioactivity and radiation exposure.

Question 1

How much aluminium is required to reduce the intensity of a 200 keV gamma-ray beam to 10% of its incident intensity? Assume that the Half Value Layer for 200 keV gamma-rays in Al is 2.14 cm.

Answer

The question phrased in terms of the symbols used above is:

$$I_x = \frac{I_0}{10} \quad , \text{ when } x = ?$$

We are told that the Half Value Layer is 2.14 cm. Therefore the Linear Attenuation Coefficient is

$$\mu = \frac{0.693}{x_{\frac{1}{2}}} = \frac{0.693}{2.14} = 0.324 \text{ cm}^{-1}$$

Now combining all this with the exponential attenuation equation:

$$I_x = I_0 \exp(-\mu x)$$

we can write:

$$\frac{I_0}{10} = I_0 \exp(-0.324x)$$

Therefore

$$\frac{1}{10} = \exp(-0.324x)$$

$$\therefore x = \frac{\ln 10}{0.324} = \frac{2.3}{0.324} = 7.1 \text{ cm}$$

$$\therefore x \approx 7 \text{ cm}$$

So the thickness of aluminium required to reduce these gamma-rays by a factor of ten is about 7 cm. This relatively large thickness is the reason why aluminium is not generally used in radiation protection - its atomic number is not high enough for efficient and significant attenuation of gamma-rays.

You might like to try this question for the case when Pb is the absorber - but you will need to find out the Half Value Layer for the 200 keV gamma-rays yourself!

Here's a hint though: have a look at one of the tables above.

And here's the answer for you to check when you've finished: **2.2 mm**.

In other words a relatively thin thickness of Pb is required to do the same job as 7 cm of aluminium.

Question 2

A 10^5 MBq source of ^{137}Cs is to be contained in a Pb box so that the exposure rate 1 m away from the source is less than 0.5 mR/hour. If the Half Value Layer for ^{137}Cs gamma-rays in Pb is 0.6 cm, what thickness of Pb is required? The Specific Gamma Ray Constant for ^{137}Cs is $3.3 \text{ R hr}^{-1} \text{ mCi}^{-1}$ at 1 cm.

Answer

This is a fairly typical question which arises when someone is using radioactive materials. We wish to use a certain quantity of the material and we wish to store it in a lead container so that the exposure rate when we are working a certain distance away is below some level for safety reasons. We know the radioactivity of the material we will be using. But its quoted in SI units. We look up a reference book to find out the exposure rate for this radioisotope and find that the Specific Gamma Ray Constant is quoted in traditional units. Just as in our question!

So let us start by getting our units right. The Specific Gamma Ray Constant is given as:

$$3.3 \text{ R hr}^{-1} \text{ mCi}^{-1} \text{ at 1 cm from the source.}$$

This is equal to:

$$3300 \text{ mR hr}^{-1} \text{ mCi}^{-1} \text{ at 1 cm from the source,}$$

which is equal to:

$$\frac{3300}{(100)^2} \text{ mR hr}^{-1} \text{ mCi}^{-1} \text{ at 1 m from the source,}$$

on the basis of the Inverse Square Law. This result expressed per becquerel is

$$\frac{3300}{10^4 (3.7 \times 10^7)} \text{ mR hr}^{-1} \text{ Bq}^{-1} \text{ at 1 m from the source,}$$

since $1 \text{ mCi} = 3.7 \times 10^7 \text{ Bq}$. And therefore for 10^5 MBq , the exposure rate is:

$$\frac{3300 \times 10^5 \times 10^6}{10^4 (3.7 \times 10^7)} \text{ mR hr}^{-1} (10^5 \text{ MBq})^{-1} \text{ at 1 m from the source,}$$

That is the exposure rate 1 meter from our source is **891.9 mR hr⁻¹**.

We wish to reduce this exposure rate according to the question to less than 0.5 mR hr^{-1} using Pb.

You should be able at this stage to use the exponential attenuation equation along with the Half Value Layer for these gamma-rays in Pb to calculate that the thickness of Pb required is about **6.5 cm**.

External Links

- [Mucal on the Web](#) - an online program which calculates x-ray absorption coefficients - by Pathikrit Bandyopadhyay, The Center for Synchrotron Radiation Research and Instrumentation at the Illinois Institute of Technology.
- [Tables of X-Ray Mass Attenuation Coefficients](#) - a vast amount of data for all elements from National Institute of Science & Technology, USA.

Multiple Choice Questions

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CHAPTER 06 • ATTENUATION OF GAMMA-RAYS

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CHAPTER 7 GAS-FILLED RADIATION DETECTORS

Introduction

We have learnt in the last two chapters about how radiation interacts with matter and we are now in a position to apply our understanding to the detection of radiation.

One of the major outcomes of the interaction of radiation with matter is the creation of ions as we saw in Chapter 5. This outcome is exploited in gas-filled detectors as you will see in this chapter. The detector in this case is essentially a gas, in that it is the atoms of a gas which are ionised by the radiation. We will see in the next chapter that solids can also be used as radiation detectors but for now we will deal with gases and be introduced to detectors such as the **Ionization Chamber** and the **Geiger Counter**.

Before considering these specific types of gas-filled detectors we will first of all consider the situation from a very general perspective.

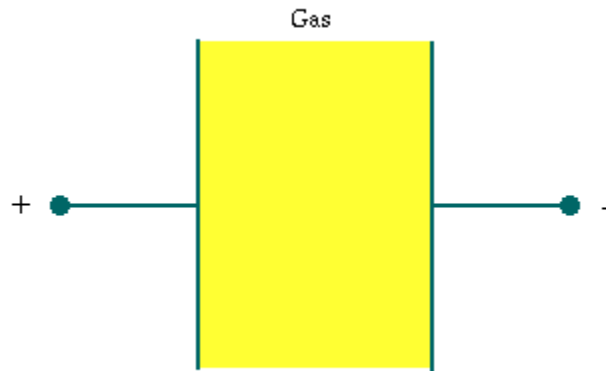
Gas-Filled Detectors

As we noted above the radiation interacts with gas atoms in this form of detector and causes ions to be produced. On the basis of what we covered in Chapter 5 it is easy to appreciate that it is the Photoelectric and Compton Effects that cause the ionisations when the radiation consists of gamma-rays with energies useful for diagnostic purposes.

There are actually two particles generated when an ion is produced - the positive ion itself and an electron. These two particles are collectively called an **ion pair**. The detection of the production of ion pairs in the gas is the basis upon which gas detectors operate. The manner in which this is done is by using an electric field to sweep the electrons away to a positively charged electrode and the ions to a negatively charged electrode.

Let us consider a very simple arrangement as shown in the following figure:

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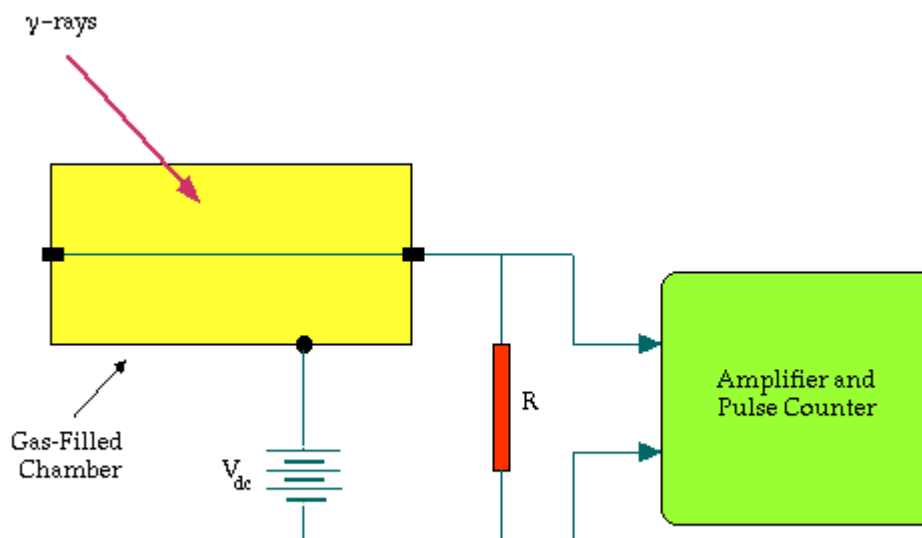
Here we have two electrodes with the gas between them. Something like a capacitor with a gas dielectric.

The gas which is used is typically an inert gas, for example **argon** or **xenon**. The reason for using an inert gas is so that chemical reactions will not occur within the gas following the ionisations which could change the characteristics of our detector.

A dc voltage is placed between the two electrodes. As a result when the radiation interacts with a gas atom the electron will move towards the positive electrode and the ion will move towards the negative electrode. But will these charges reach their respective electrodes? The answer is obviously dependent on the magnitude of the dc voltage. For example if at one extreme we had a dc voltage of a microvolt (that is, one millionth of a volt) the resultant electric field may be insufficient to move the ion pair very far and the two particles may recombine to reform the gas atom. At the other extreme suppose we applied a million volts between the two electrodes. In this case we are likely to get sparks flying between the two electrodes - a lightning bolt if you like - and our detector might act something like a **neon sign**. Somewhere in between these two extremes though we should be able to provide a sufficient attractive force for the ion and electron to move to their respective electrodes without recombination or sparking occurring.

We will look at this subject in more detail below. Before we do let us see how the concept of the simple detector illustrated above is applied in practice. The gas-filled chamber is generally cylindrical in shape in real detectors. This shape has been found to be more efficient than the parallel electrode arrangement shown above.

A cross-sectional view through this cylinder is shown in the following figure:



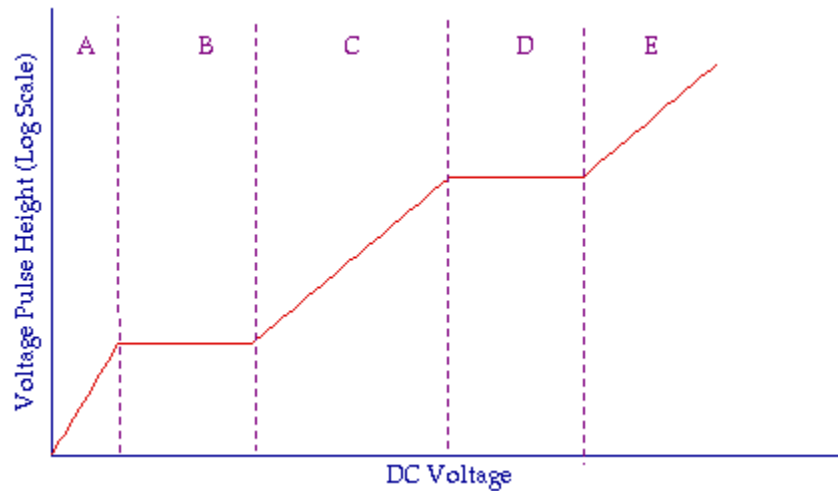
The positive electrode consists of a thin wire running through the centre of the cylinder and the negative electrode consists of the wall of the cylinder. In principle we could make such a detector by getting a section of a metal pipe, mounting a wire through its centre, filling it with an inert gas and sealing the ends of the pipe. Actual detectors are a little bit more complex however but let us not get side-tracked at this stage.

We apply a dc voltage via a battery or via a dc voltage supply and connect it as shown in the figure using a resistor, R . Now, assume that a gamma-ray enters the detector. Ion pairs will be produced in the gas - the ions heading towards the outer wall and the electrons heading towards the centre wire. Let us think about the electrons for a moment. When they hit the centre wire we can simply think of them as entering the wire and flowing through the resistor to get to the positive terminal of the dc voltage supply. These electrons flowing through the resistor constitute an electric current and as a result of **Ohm's Law** a voltage is generated across the resistor. This voltage is amplified by an amplifier and some type of device is used to register the amplified voltage. A loud-speaker is a fairly simple device to use for this purpose and the generation of a voltage pulse is manifest by a click from the loud-speaker. Other display devices include a **ratemeter** which displays the number of voltage pulses generated per unit time - something like a speedometer in a car - and a pulse counter (or **scaler**) which counts the number of voltage pulses generated in a set period of time. A voltage pulse is frequently referred to in practice as a **count** and the number of voltage pulses generated per unit time is frequently called the **count rate**.

DC Voltage Dependence

If we were to build a detector and electronic circuit as shown in the figure above we could conduct an experiment that would allow us to explore the effect of the dc voltage on the magnitude of the voltage pulses produced across the resistor, R . Note that the term **pulse height** is frequently used in this field to refer to the magnitude of voltage pulses.

Ideally, we could generate a result similar to that illustrated in the following figure:



The graph illustrates the dependence of the pulse height on the dc voltage. Note that the vertical axis representing the pulse height is on a logarithmic scale for the sake of compressing a large linear scale onto a reasonably-sized graph.

The experimental results can be divided into five regions as shown. We will now consider each region in turn.

- **Region A** Here V_{dc} is relatively low so that recombination of positive ions and electrons occurs. As a result not all ion pairs are collected and the voltage pulse height is relatively low. It does increase as the dc voltage increases however as the amount of recombination reduces.
- **Region B** V_{dc} is sufficiently high in this region so that only a negligible amount of recombination occurs. This is the region where a type of detector called the **Ionization Chamber** operates.
- **Region C** V_{dc} is sufficiently high in this region so that electrons approaching the centre wire attain sufficient energy between collisions with the electrons of gas atoms to produce new ion pairs. Thus the number of electrons is increased so that the electric charge passing through the resistor, R , may be up to a thousand times greater than the charge produced initially by the radiation interaction. This is the region where a type of detector called the **Proportional Counter** operates.
- **Region D** V_{dc} is so high that even a minimally-ionizing particle will produce a very large voltage pulse. The initial ionization produced by the radiation triggers a complete gas breakdown as an avalanche of electrons heads towards and spreads along the centre wire. This region is called the **Geiger-Müller Region**, and is exploited in the Geiger Counter.
- **Region E** Here V_{dc} is high enough for the gas to completely breakdown and it cannot be used to detect radiation.

We will now consider features of the Ionisation Chamber and the Geiger Counter in more detail.

Ionisation Chamber

The ionisation chamber consists of a gas-filled detector energised by a relatively low dc voltage. We will first of all make an estimate of the voltage pulse height generated by this type of detector. We

will then consider some applications of ionisation chambers.

When a beta-particle interacts with the gas the energy required to produce one ion pair is about 30 eV. Therefore when a beta-particle of energy 1 MeV is completely absorbed in the gas the number of ion pairs produced is:

$$n = \frac{1 \text{ MeV}}{30 \text{ eV}} = \frac{1 \times 10^6}{30} \approx 3 \times 10^4 \text{ ion pairs}$$

The electric charge produced in the gas is therefore

$$Q = n.e$$

$$\therefore Q = (3 \times 10^4 \text{ ion pairs}).(1.6 \times 10^{-19} \text{ C})$$

$$\therefore Q = 5 \times 10^{-15} \text{ C}$$

If the capacitance of the ionisation chamber (remember that we compared a gas-filled detector to a capacitor above) is 100 pF then the amplitude of the voltage pulse generated is:

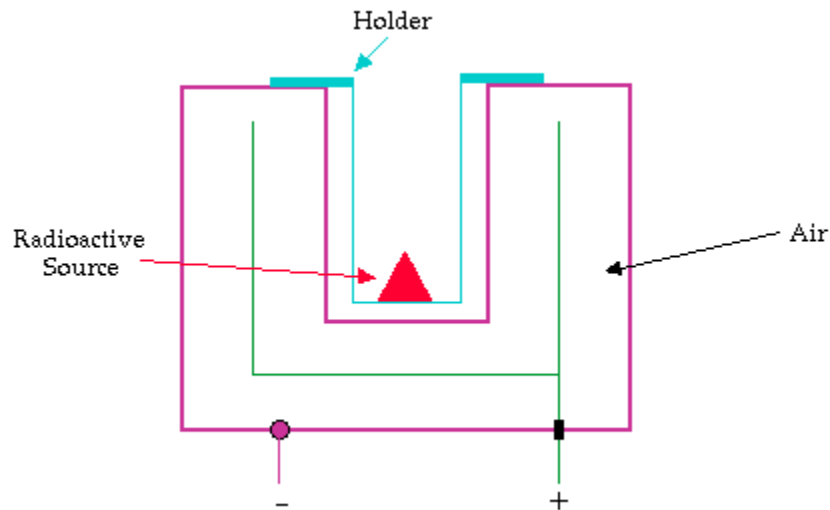
$$V = \frac{Q}{C} = \frac{5 \times 10^{-15} \text{ C}}{100 \times 10^{-12} \text{ F}} = 5 \times 10^{-5} \text{ V}$$

$$\therefore V = 50 \mu\text{V}$$

Because such a small voltage is generated it is necessary to use a very sensitive amplifier in the electronic circuitry connected to the chamber.

We will now learn about two applications of ionisation chambers. The first one is for the measurement of radiation exposures. You will remember from [Chapter 4](#) that the unit of radiation exposure (be it the SI or the traditional unit) is defined in terms of the amount of electric charge produced in a unit mass of a air. An ionization chamber filled with air is the natural instrument to use for such measurements.

The second application is the measurement of radioactivity. The ionisation chamber used here is configured in what is called a **re-entrant** arrangement (see figure below) so that the sample of radioactive material can be placed within the detector using a holder and hence most of the emitted radiation can be detected. The instrument is widely referred to as an **Isotope Calibrator** and the trickle of electric current generated by such a detector is calibrated so that a reading in units of radioactivity (for example MBq or mCi) can be obtained. Most well-run Nuclear Medicine Departments will have at least one of these devices so that doses of radioactivity can be checked prior to administration to patients.



Here are some photographs of ionisation chambers designed for various applications:



An exposure-area product detector used in radiography.



A range of ionisation chambers of different volumes used for measuring radiation exposure.



An exposure meter used in radiography.



An isotope calibrator used in nuclear medicine - the blue cylinder on the left contains the re-entrant chamber.



An exposure meter used in radiography.

Geiger Counter

We saw earlier that the Geiger Counter operates at relatively high dc voltages (for example 300-400 volts) and that an avalanche of electrons is generated following the absorption of radiation in the

gas. The voltage pulses produced by this detector are relatively large since the gas effectively acts as an amplifier of the electric charge produced.

There are four features of this detector which we will discuss. The first is that a sensitive amplifier (as was the case with the Ionization Chamber) is not required for this detector because of the gas amplification noted above.

The second feature results from the fact that the generation of the electron avalanche must be stopped in order to reform the detector. In other words when a radiation particle/photon is absorbed by the gas a complete gas breakdown occurs which implies that the gas is incapable of detecting the next particle/photon which enters the detector. So in the extreme case one minute we have a radiation detector and the following moment we do not.

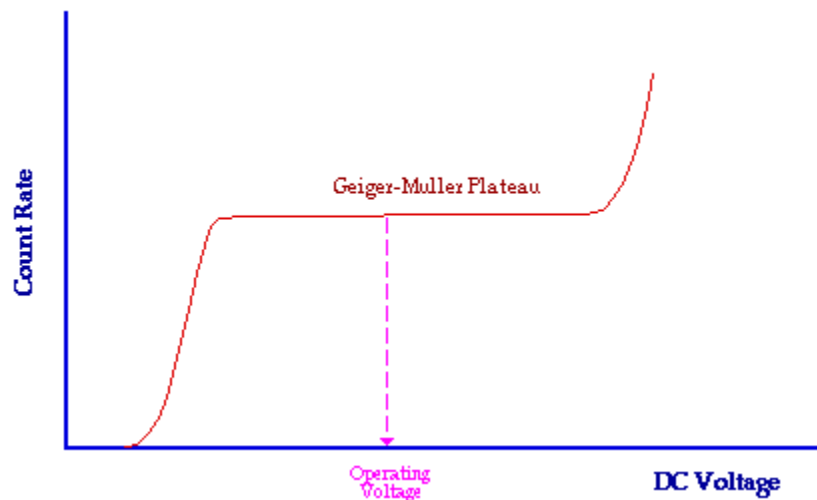
A means of stopping the electron avalanche is therefore required - a process called **Quenching**. One means of doing this is by electronically lowering the dc voltage following an avalanche. A more widely used method of quenching is to add a small amount of a **quenching gas** to the inert gas. For example the gas could be argon with **ethyl alcohol** added. The ethyl alcohol is in vapour form and since it consists of relatively large molecules energy which would in their absence give rise to sustaining the electron avalanche is absorbed by these molecules. The large molecules act like a brake in effect.

Irrespective of the type of quenching used the detector is insensitive for a small period of time following absorption of a radiation particle/photon. This period of time is called the **Dead Time** and this is the third feature of this detector which we will consider. Dead times are relatively short but nevertheless significant - being typically of the order of 200-400 μs . As a result the reading obtained with this detector is less than it should be. The true reading without going into detail can be obtained using the following equation:

$$T = \frac{A}{1 - \tau A}$$

where T is the true reading, A is the actual reading and τ is the dead time. Some instruments perform this calculation automatically.

The fourth feature to note about this detector is the dependence of its performance on the dc voltage. The Geiger-Müller Region of our figure above is shown in more detail below:



Notice that it contains a plateau where the count rate obtained is independent of the dc voltage. The centre of this plateau is where most detectors are operated. It is clear that the count rate from the detector is not affected if the dc voltage fluctuates about the operating voltage. This implies that a relatively straight-forward dc voltage supply can be used. This feature coupled with the fact that a sensitive amplifier is not needed translates in practice to a relatively inexpensive radiation detector.

External Links

- [Inside a smoke detector](#) - about the ion chamber used in smoke detectors - from the How Stuff Works website.
- [Ionisation Chambers](#) - a brief description from the Triumph Safety Group.
- [Radiation and Radioactivity](#) - a self-paced lesson developed by the University of Michigan's Student Chapter of the Health Physics Society with a section on gas filled detectors.
- [The Geiger Counter](#) - a brief overview from the NASA Goddard Space Flight Center, USA.

Multiple Choice Questions

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CHAPTER 8 SCINTILLATION DETECTORS

Introduction

The second type of radiation detector we will discuss is called the scintillation detector.

Scintillations are minute flashes of light which are produced by certain materials when they absorb radiation. These materials are variously called fluorescent materials, flours, **scintillators** or phosphors.

If we had a radioactive source and a scintillator in the lab we could darken the room, move the scintillator close to the source and see the scintillations. These small flashes of light might be green or blue or some other colour depending on the scintillator. We could also count the number of flashes produced to gain an estimate of the radioactivity of the source, that is the more flashes of light seen the more radiation present.

The scintillation detector was possibly the first radiation detector discovered. You might have heard the story of the discovery of X-rays by **Wilhelm Roentgen** in 1895. He was working one evening in his laboratory in Wurzburg, Germany with a device which fired a beam of electrons at a target inside an evacuated glass tube. While working with this device he noticed that some platino-barium cyanide crystals, which he just happened to have close by, began to glow - and that they stopped glowing when he switched the device off. Roentgen had accidentally discovered a new form of radiation. He had also accidentally discovered a scintillator detector.

Although scintillations can be seen we have a more sophisticated way of counting and measuring them today by using some form of **photodetector**.

We will learn about the construction and mode of operation of this type of detector in this chapter. In addition, we will see how it can be used not just for detecting the presence of ionizing radiation but also for measuring the **energy** of that radiation.

Before we do however it is useful to note that scintillators are very widely used in the medical radiations field. For example the **X-ray cassette** used in radiography contains a scintillator (called an **intensifying screen**) in close contact with a photographic film. A second example is the **X-ray Image Intensifier** used in **fluoroscopy** which contains scintillators called phosphors. Scintillators are also used in some **CT Scanners** and as we will see in the **next chapter**, in the **Gamma Camera** and **PET Scanner**. Their application is not limited to the medical radiations field in that scintillators are also used as screens in **television sets** and **computer monitors** and for generating light in **fluorescent tubes** - to mention just two common applications. What other applications can you think of?

So scintillators are a lot more common than you might initially think and you will therefore find the information presented here useful to you not just for your studies of nuclear medicine.

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Fluorescent Materials

Some fluorescent materials are listed in the following table. Thallium-activated **sodium iodide**, NaI(Tl) is a crystalline material which is widely used for the detection of gamma-rays in scintillation detectors. We will be looking at this in more detail below.

Material	Form
NaI (Tl)	crystal
CsI (Na)	crystal
CaWO ₄	crystal
ZnS (Ag)	powder
p-terphenyl in toluene	liquid
p-terphenyl in polystyrene	plastic

Another crystalline material sodium-activated **caesium iodide**, CsI(Na) is widely used for X-ray detection in devices such as the **X-ray image intensifier**. Another one called calcium tungstate, CaWO₄ has been widely used in X-ray cassettes although this substance has been replaced by other scintillators such as lanthanum oxybromide in many modern cassettes.

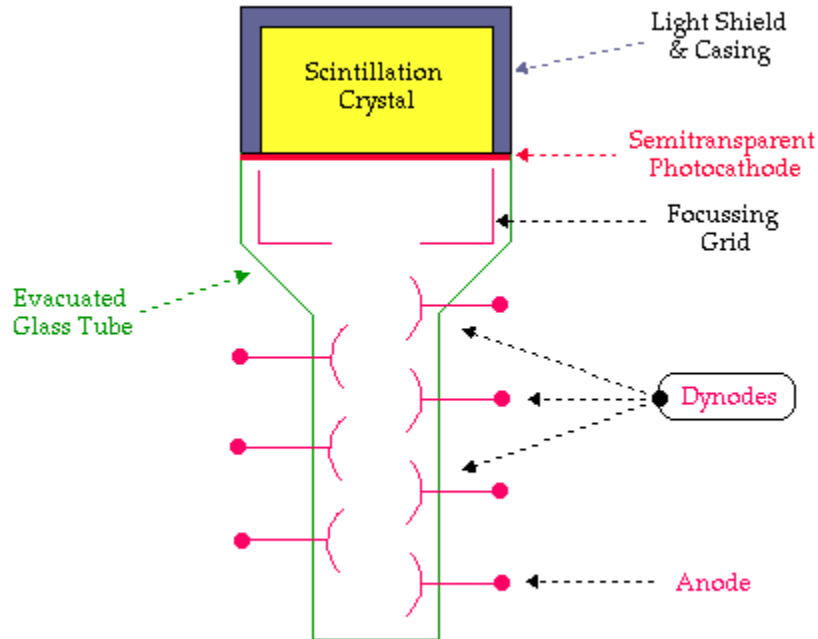
Notice that some scintillation materials are activated with certain elements. What this means is that the base material has a small amount of the activation element present. The term **doped** is sometimes used instead of **activated**. This activating element is used to influence the wavelength (colour) of the light produced by the scintillator.

Silver-activated **zinc sulphide** is a scintillator in powder form and p-terphenyl in toluene is a **liquid scintillator**. The advantage of such forms of scintillators is that the radioactive material can be placed in close contact with the radioactive material. For example if a radioactive sample happened to be in liquid form we could mix it with a liquid scintillator so as to optimise the chances of detection of the emitted radiation and hence have a very sensitive detector.

A final example is p-terphenyl in polystyrene which is a scintillator in the form of a plastic. This form can be easily made into different shapes like most plastics and is therefore useful when detectors of particular shapes are required.

Photomultiplier Tube

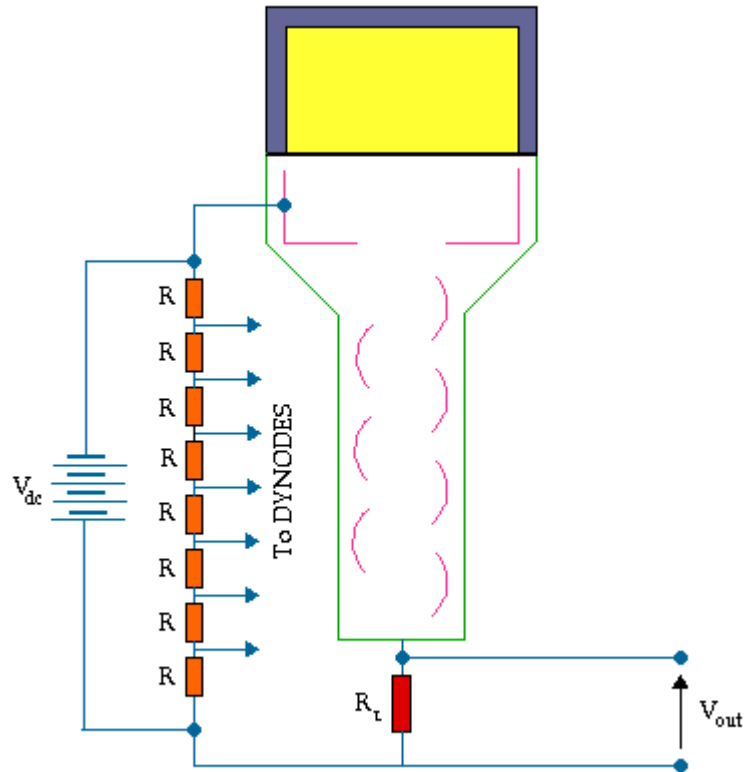
A scintillator coupled to a **photomultiplier tube** (PMT) is illustrated in the following figure. The overall device is typically cylindrical in shape and the figure shows a cross-section through this cylinder:



The scintillation crystal, NaI(Tl) is very delicate and this is one of the reasons it is housed in an aluminium casing. The inside wall of the casing is designed so that any light which strikes it is reflected downwards towards the PMT.

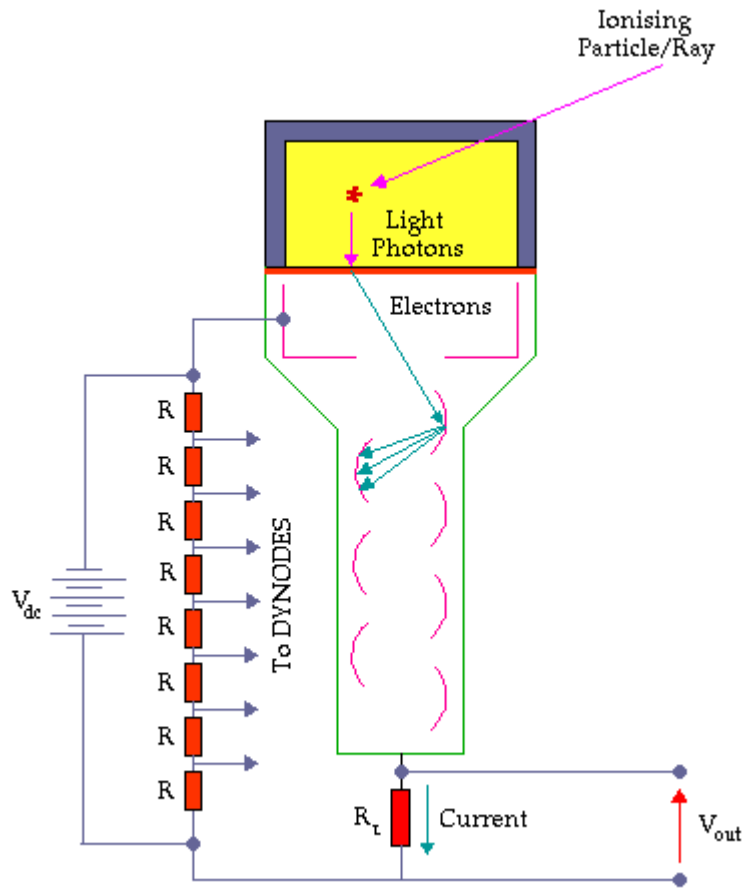
The PMT itself consists of a photocathode, a focussing grid, an array of dynodes and an anode housed in an evacuated glass tube. The function of the photocathode is to convert the light flashes produced by radiation attenuation in the scintillation crystal into electrons. The grid focuses these electrons onto the first dynode and the dynode array is used for electron multiplication. We will consider this process in more detail below. Finally the anode collects the electrons produced by the array of dynodes.

The electrical circuitry which is typically attached to a PMT is shown in the next figure:



It consists of a high voltage supply, a resistor divider chain and a load resistor, R_L . The high voltage supply generates a dc voltage, V_{dc} which can be up to 1,000 volts. It is applied to the resistor divider chain which consists of an array of resistors, each of which has the same resistance, R . The function of this chain of resistors is to divide up V_{dc} into equal voltages which are supplied to the dynodes. As a result voltages which increase in equal steps are applied to the array of dynodes. The load resistor is used so that an output voltage, V_{out} can be generated.

Finally the operation of the device is illustrated in the figure below:



The ionizing radiation produces flashes of light in the scintillation crystal. This light strikes the photocathode and is converted into electrons. The electrons are directed by the grid onto the first dynode.

Dynodes are made from certain alloys which emit electrons when their surface is struck by electrons with the advantage that more electrons are emitted than are absorbed. A dynode used in a PMT typically emits between two and five electrons for each electron which strikes it.

So when an electron from the photocathode strikes the first dynode between two and five electrons are emitted and are directed towards the second dynode in the array (three are illustrated in the figure). This electron multiplication process is repeated at the second dynode so that we end up with nine electrons for example heading towards the third dynode. An electron avalanche therefore develops so that a sizeable number of electrons eventually hits the anode at the bottom of the dynode chain.

These electrons flow through the load resistor, R_L and constitute an electric current which according to Ohm's Law generates a voltage, V_{out} which is measured by electronic circuitry (which we will describe later).

An animated graphic of the operation of the scintillation detector can be viewed by clicking [HERE](#).

A number of photographs of devices based on scintillation detection are shown below:



An X-ray cassette used in radiography.



A NaI crystal coupled to a photomultiplier tube.



A number of identical photomultiplier tubes from a Gamma Camera.



A single channel analyser.



A single channel analyser with a crystal-PMT assembly.



A well detector.

The **important feature** of the scintillation detector is that this output voltage, V_{out} is directly proportional to the energy deposited by the radiation in the crystal. We will see what a useful feature this is below. Before we do so we will briefly analyze the operation of this device.

Mathematical Model

A simple mathematical model will be presented below which will help us get a better handle on the performance of a scintillation detector. We will do this by quantifying the performance of the scintillator, the photocathode and the dynodes.

Let's use the following symbols to characterize each stage of the detection process:

- m : number of light photons produced in crystal
- k : optical efficiency of the crystal, that is the efficiency with which the crystal transmits light
- l : **quantum efficiency** of the photocathode, that is the efficiency with which the photocathode converts light photons to electrons
- n : number of dynodes
- R : dynode multiplication factor, that is the number of secondary electrons emitted by a

dynode per primary electron absorbed.

Therefore the charge collected at the anode is given by the following equation:

$$Q = m k l R^n e \quad \text{where } e: \text{ the electronic charge.}$$

For example supposing a 100 keV gamma-ray is absorbed in the crystal. The number of light photons produced, m , might be about 1,000 for a typical scintillation crystal. A typical crystal might have an optical efficiency, k , of 0.5 - in other words 50% of the light produced reaches the photocathode which might have a quantum efficiency of 0.15. A typical PMT has ten dynodes and let us assume that the dynode multiplication factor is 4.5.

Therefore

$$Q = 1000(0.5)(0.15)(4.5^{10})(1.6 \times 10^{-19}) \text{ C}$$

$$\therefore Q = 41 \times 10^{-12} \text{ C}$$

$$\therefore Q \approx 40 \text{ pC}$$

This amount of charge is very small. Even though we have used a sophisticated photodetector like a PMT we still end up with quite a small electrical signal.

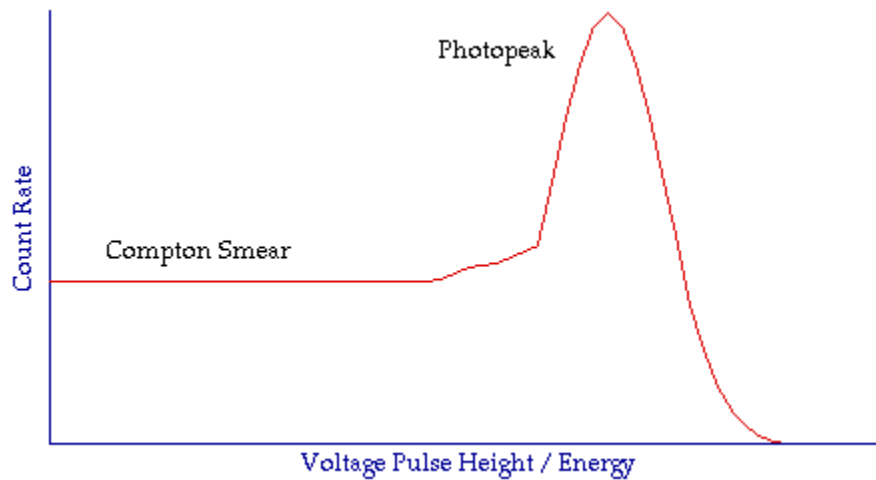
A very sensitive amplifier is therefore needed to amplify this signal. This type of amplifier is generally called a **pre-amplifier** and we will refer to it again later.

Output Voltage

We noted above that the voltage measured across the resistor, R_L , is proportional to the energy deposited in the scintillation crystal by the radiation. Let us consider how the radiation might deposit its energy in the crystal.

Let us consider a situation where gamma-rays are detected by the crystal. We learnt in [Chapter 5](#) that there were two interaction mechanisms involved in gamma-ray attenuation - the Photoelectric Effect and the Compton Effect. You will remember that the Photoelectric Effect involves the total absorption of the energy of a gamma-ray, while the Compton Effect involves just partial absorption of this energy. Since the output voltage of a scintillation detector is proportional to the energy deposited by the gamma-rays it is reasonable to expect that Photoelectric Effects in the crystal will generate distinct and relatively large output voltages and that Compton Effects will result in lower output voltages.

The usual way of presenting this information is by plotting a graph of the count rate versus the output voltage pulse height as shown in the following figure:



This plot illustrates what is obtained for a monoenergetic gamma-emitting radioisotope, for example ^{99m}Tc - which, as we have noted before emits a single gamma-ray with an energy of 140 keV.

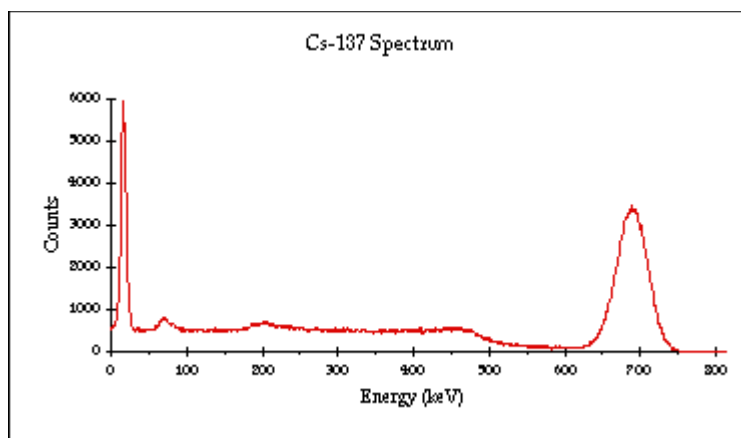
Before we look at it in detail remember that we noted above that the output voltage from this detector is proportional to the energy deposited by the radiation in the crystal. The horizontal axis can therefore be used to represent the output voltage or the gamma-ray energy. Both of these quantities are shown in the figure to help with this discussion. In addition note that this plot is often called a **Gamma-Ray Energy Spectrum**.

The figure above contains two regions. One called the **Photopeak** and the other called the **Compton Smear**. The Photopeak results because of Photoelectric absorption of the gamma-rays from the radioactive source - remember that we are dealing with a monoenergetic emitter in this example. It consists of a peak representing the gamma-ray energy (140 keV in our example). If our radioisotope emitted gamma-rays of two energies we would have two photopeaks in our spectrum and so on.

Notice that the peak has a statistical spread. This has to do with how good our detector is and we will not get into any detail about it here other than to note that the extent of this spread is a measure of the quality of our detector. A high quality (and more expensive!) detector will have a narrower statistical spread in the photopeaks which it measures.

The other component of our spectrum is the Compton Smear. It represents a range of output voltages which are lower than that for the Photopeak. It is therefore indicative of the partial absorption of the energy of gamma-rays in the crystal. In some Compton Effects a substantial scattering with a valence electron can occur which gives rise to relatively large voltage pulses. In other Compton Effects the gamma-ray just grazes off a valence electron with minimal energy transfer and hence a relatively small voltage pulse is generated. In between these two extremes are a range of scattering events involving a range of energy transfers and hence a range of voltage pulse heights. A 'smear' therefore manifests itself on the gamma-ray energy spectrum.

It is important to note that the spectrum illustrated in the figure is simplified for the sake of this introductory discussion and that actual spectra are a little more complex - see figure below for an example:



You will find though that your understanding of actual spectra can easily develop on the basis of the simple picture we have painted here.

It is also important to appreciate the additional information which this type of radiation detector provides relative to a gas-filled detector. In essence gas-filled detectors can be used to tell us if any radiation is present as well as the amount of that radiation. Scintillation detectors also give us this information but they tell us about the **energy** of this radiation as well. This additional information can be used for many diverse applications such as the identification of unknown radioisotopes and the production of nuclear medicine images. Let us stay a little bit longer though with the fundamental features of how scintillation detectors work.

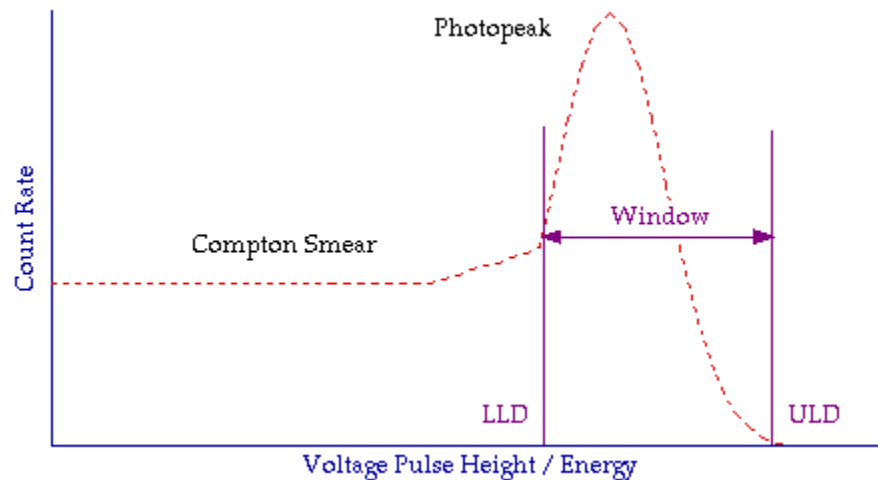
The photopeak of the Gamma-Ray Energy Spectrum is generally of interest in nuclear medicine. This peak is the main signature of the radioisotope being used and its isolation from the Compton Smear is normally achieved using a technique called **Pulse Height Analysis**.

Pulse Height Analysis

This is an electronic technique which allows a spectrum to be acquired using two types of circuitry. One circuit is called a **Lower Level Discriminator** which only allows voltages pulses through it which are lower than its setting. The other is called an **Upper Level Discriminator** which only allows voltage pulses though which are (you guessed it!) higher than its setting.

The result of using both these circuits in combination is a variable-width window which can be placed anywhere along a spectrum. For example if we wished to obtain information from the photopeak only of our simplified spectrum we would place the discrimination controls as shown in the following figure:

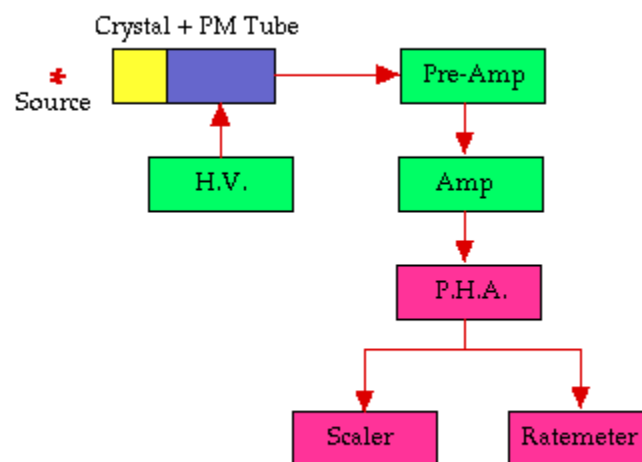
A final point to note here is that since the scintillation detector is widely used to obtain information about the energies of the radiation emitted from a radioactive source it is frequently referred to as a **Scintillation Spectrometer**.



Scintillation Spectrometer

Types of scintillation spectrometer fall into two basic categories - the relatively straight-forward **Single Channel Analyser** and the more sophisticated **Multi-Channel Analyser**.

The Single Channel Analyser is the type of instrument we have been describing so far in this discussion. A block diagram of the instrument is shown below:

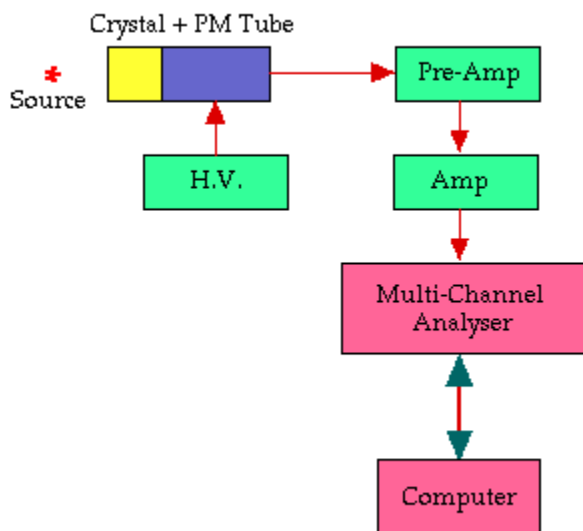


It consists of a scintillation crystal coupled to a photomultiplier tube which is powered by a high-voltage circuit (H.V.). The output voltages are initially amplified by a sensitive pre-amplifier (**Pre-Amp**) as we noted above before being amplified further and conditioned by the amplifier (**Amp**).

The voltage pulses are then in a suitable form for the pulse height analyser (**P.H.A.**) - the output pulses from which can be fed to a **Scaler** and a **Ratemeter** for display of the information about the portion of the spectrum we have allowed to pass through the PHA. The Ratemeter is a display device just like the speedometer in a car and indicates the number of pulses generated per unit time. The Scaler on the other hand usually consists of a digital display which shows the number of voltage pulses produced in a specified period of time.

We can illustrate the operation of this circuitry by considering how it might be used to generate a Gamma-Ray Energy Spectrum. What we would do is set up the LLD and ULD so as to define a narrow window and place this to pass the lowest voltage pulses produced by the detector through to the Scaler and Ratemeter. In other words we would place a narrow window at the extreme left of the spectrum and acquire information about the lowest energy gamma-ray interactions in the crystal. We would then adjust the LLD and ULD settings to acquire information about the interactions of the next highest energy. We would proceed in this fashion to scan the whole spectrum.

A more sophisticated detector circuit is illustrated in the following figure:



It is quite similar to that in the previous figure with the exception that the PHA, Scaler and Ratemeter are replaced by a Multi-Channel Analyser and a computer. The Multi-Channel Analyser (MCA) is a circuit which is capable of setting up a large number of individual windows to look at a complete spectrum in one go. The MCA might consist of 1024 individual windows for example and the computer might consist of a personal computer which can acquire information simultaneously from each window and display it as an energy spectrum. The computer generally contains software which allows us to manipulate the resultant information in a variety of ways. Indeed the ^{137}Cs spectrum shown above was generated using this approach.

External Links

- [Radiation and Radioactivity](#) - a self-paced lesson developed by the University of Michigan's Student Chapter of the Health Physics Society, with a section on sodium iodide detectors.

Multiple Choice Questions

Click [HERE](#) to access an online MCQ covering the material in this chapter.

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CHAPTER 9 NUCLEAR MEDICINE IMAGING SYSTEMS

Introduction

Topics we have covered in this wikibook have included radioactivity, the interaction of gamma-rays with matter and radiation detection. The main reason for following this pathway was to bring us to the subject of this chapter: nuclear medicine imaging systems. These are devices which produce pictures of the distribution of radioactive material following administration to a patient.

The radioactivity is generally administered to the patient in the form of a **radiopharmaceutical** - the term **radiotracer** is also used. This follows some physiological pathway to accumulate for a short period of time in some part of the body. A good example is ^{99m}Tc -tin colloid which following intravenous injection accumulates mainly in the patient's liver. The substance emits gamma-rays while it is in the patient's liver and we can produce an image of its distribution using a nuclear medicine imaging system. This image can tell us whether the function of the liver is normal or abnormal or if sections of it are damaged from some form of disease.

Different radiopharmaceuticals are used to produce images from almost all regions of the body:

Part of the Body	Example Radiotracer
Brain	^{99m}Tc -Ceretec
Thyroid	$\text{Na}^{99m}\text{TcO}_4$
Lung (Ventilation)	^{133}Xe gas
Lung (Perfusion)	^{99m}Tc -MAA
Liver	^{99m}Tc -Tin Colloid
Spleen	^{99m}Tc -Damaged Red Blood Cells
Pancreas	^{75}Se -Selenomethionine
Kidneys	^{99m}Tc -DMSA

Note that the form of information obtained using this imaging method is mainly related to the physiological functioning of an organ as opposed to the mainly anatomical information which is obtained using X-ray imaging systems. Nuclear medicine therefore provides a different perspective on

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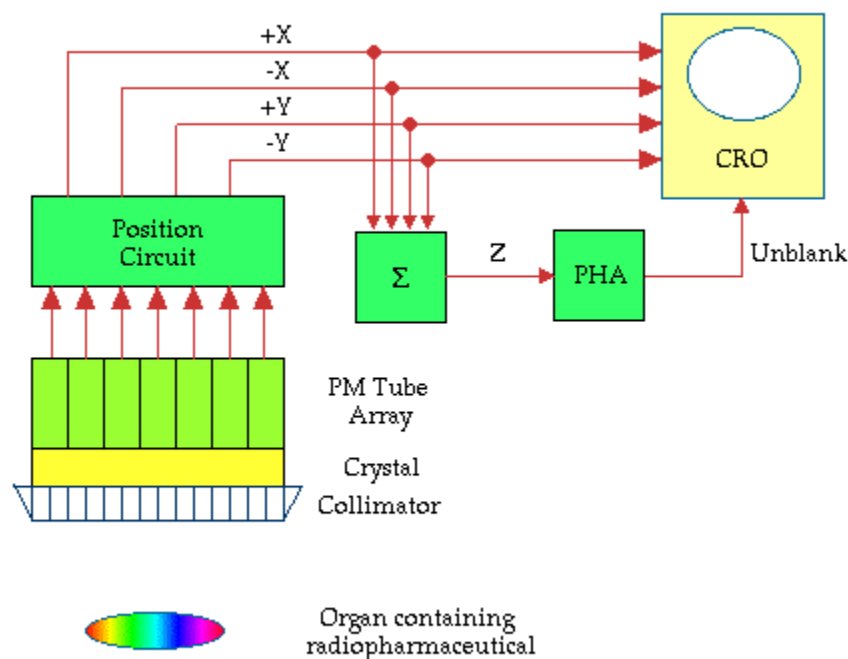
a disease condition and generates additional information to that obtained from X-ray images. Our purpose here is to concentrate on the imaging systems used to produce the images.

Early forms of imaging system used in this field consisted of a radiation detector (a scintillation detector for example) which was scanned slowly over a region of the patient in order to measure the radiation intensity emitted from individual points within the region. One such device was called the **Rectilinear Scanner**. Such imaging systems have been replaced since the 1970s by more sophisticated devices which produce images much more rapidly. The most common of these modern devices is called the **Gamma Camera** and we will consider its construction and mode of operation below.

Gamma Camera

The basic design of the most common type of gamma camera used today was developed by an American physicist, **Hal Anger** and is therefore sometimes called the Anger Camera. It consists of a large diameter NaI(Tl) scintillation crystal which is viewed by a large number of photomultiplier tubes.

A block diagram of the basic components of a gamma camera is shown below:



The crystal and PM Tubes are housed in a cylindrical shaped housing commonly called the **camera head** and a cross-sectional view of this is shown in the figure. The crystal can be between about 25 cm and 40 cm in diameter and about 1 cm thick. The diameter is dependent on the application of the device. For example a 25 cm diameter crystal might be used for a camera designed for cardiac applications while a larger 40 cm crystal would be used for producing images of the lungs. The thickness of the crystal is chosen so that it provides good detection for the 140 keV gamma-rays emitted from ^{99m}Tc - which is the most common radioisotope used today.

Scintillations produced in the crystal are detected by a large number of PM tubes which are

arranged in a two-dimensional array. There is typically between 37 and 91 PM tubes in modern gamma cameras. The output voltages generated by these PM tubes are fed to a position circuit which produces four output signals called $\pm X$ and $\pm Y$. These position signals contain information about where the scintillations were produced within the crystal. In the most basic gamma camera design they are fed to a **cathode ray oscilloscope (CRO)**. We will describe the operation of the CRO in more detail below.

Before we do so we should note that the position signals also contain information about the intensity of each scintillation. This intensity information can be derived from the position signals by feeding them to a summation circuit (marked Σ in the figure) which adds up the four position signals to generate a voltage pulse which represents the intensity of a scintillation. This voltage pulse is commonly called the **Z-pulse** (or zee-pulse in American English!) which following pulse height analysis (**PHA**) is fed as the **unblank pulse** to the CRO.

So we end up with four position signals and an unblank pulse sent to the CRO. Let us briefly review the operation of a CRO before we continue. The core of a CRO consists of an evacuated tube with an electron gun at one end and a phosphor-coated screen at the other end. The electron gun generates an electron beam which is directed at the screen and the screen emits light at those points struck by the electron beam. The position of the electron beam can be controlled by vertical and horizontal deflection plates and with the appropriate voltages fed to these plates the electron beam can be positioned at any point on the screen. The normal mode of operation of an oscilloscope is for the electron beam to remain switched on. In the case of the gamma camera the electron beam of the CRO is normally switched off - it is said to be **blanked**.

When an unblank pulse is generated by the PHA circuit the electron beam of the CRO is switched on for a brief period of time so as to display a flash of light on the screen. In other words the voltage pulse from the PHA circuit is used to unblank the electron beam of the CRO.

So where does this flash of light occur on the screen of the CRO? The position of the flash of light is dictated by the $\pm X$ and $\pm Y$ signals generated by the position circuit. These signals as you might have guessed are fed to the deflection plates of the CRO so as to cause the unblanked electron beam to strike the screen at a point related to where the scintillation was originally produced in the NaI(Tl) crystal. Simple!

The gamma camera can therefore be considered to be a sophisticated arrangement of electronic circuits used to translate the position of a flash of light in a scintillation crystal to a flash of light at a related point on the screen of an oscilloscope. In addition the use of a pulse height analyser in the circuitry allows us to translate the scintillations related only to photoelectric events in the crystal by rejecting all voltage pulses except those occurring within the photopeak of the gamma-ray energy spectrum.

Let us **summarise** where we have got to before we proceed. A radiopharmaceutical is administered to the patient and it accumulates in the organ of interest. Gamma-rays are emitted in all directions from the organ and those heading in the direction of the gamma camera enter the crystal and produce scintillations (note that there is a device in front of the crystal called a **collimator** which we will discuss later). The scintillations are detected by an array of PM tubes whose outputs are fed to a position circuit which generates four voltage pulses related to the position of a scintillation within the crystal. These voltage pulses are fed to the deflection circuitry of the CRO. They are also fed to a summation circuit whose output (the Z-pulse) is fed to the PHA and the output of the PHA is used to switch on (that is, unblank) the electron beam of the CRO. A flash of light appears on the screen of the

CRO at a point related to where the scintillation occurred within the NaI(Tl) crystal. An image of the distribution of the radiopharmaceutical within the organ is therefore formed on the screen of the CRO when the gamma-rays emitted from the organ are detected by the crystal.

What we have described above is the operation of a fairly traditional gamma camera. Modern designs are a good deal more complex but the basic design has remained much the same as has been described. One area where major design improvements have occurred is the area of image formation and display. The most basic approach to image formation is to photograph the screen of the CRO over a period of time to allow integration of the light flashes to form an image on **photographic film**. A stage up from this is to use a **storage oscilloscope** which allows each flash of light to remain on the screen for a reasonable period of time.

The most modern approach is to feed the position signals into the memory circuitry of a computer for storage. The memory contents can therefore be displayed on a computer monitor and can also be manipulated (that is **processed**) in many ways. For example various colours can be used to represent different concentrations of a radiopharmaceutical within an organ. Click **HERE** to access an online demonstration of such digital image processing.

The use of such **digital image processing** is now widespread in nuclear medicine in that it can be used to rapidly and conveniently control image acquisition and display as well as to analyse an image or sequences of images, to annotate images with the patient's name and examination details, to store the images for subsequent retrieval and to communicate the image data to other computers over a **network**.

Some photographs of gamma cameras and related devices are shown below:



A single-headed gamma camera.



Another single-headed gamma camera.



The NaI crystal of a gamma camera.



The cathode ray oscilloscope (CRO) of a gamma camera.



The image processing system of a gamma camera.



A dual-headed gamma camera.



Another view of a dual-headed gamma camera.



The image acquisition and processing console of a dual-headed gamma camera.

We will continue with our description of the gamma camera by considering the construction and purpose of the collimator.

Collimation

The **collimator** is a device which is attached to the front of the gamma camera head. It functions something like a lens used in a photographic camera but this analogy is not quite correct because it is rather difficult to focus gamma-rays. Nevertheless in its simplest form it is used to block out all gamma rays which are heading towards the crystal except those which are travelling at right angles to the plane of the crystal:

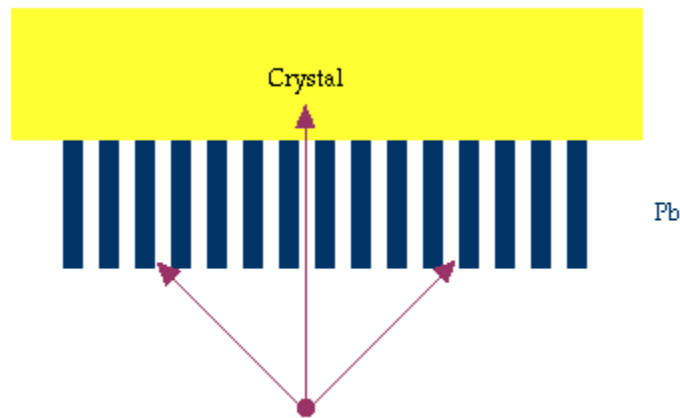


Illustration 2: Diagram of parallel-hole collimator attached to a crystal of a gamma camera. Obliquely incident gamma-rays are absorbed by the septa.

The figure illustrates a magnified view of a **parallel-hole collimator** attached to a crystal. The collimator simply consists of a large number of small holes drilled in a lead plate. Notice that gamma-rays entering at an angle to the crystal get absorbed by the lead and that only those entering along the direction of the holes get through to cause scintillations in the crystal. If the collimator was not in place these obliquely incident gamma-rays would blur the images produced by the gamma camera. In other words the images would not be very clear.

Most gamma cameras have a number of collimators which can be fitted depending on the examination. The basic design of these collimators is the same except that they vary in terms of the diameter of each hole, the depth of each hole and the thickness of lead between each hole (commonly called the **septum thickness**). The choice of a specific collimator is dependent on the amount of radiation absorption that occurs (which influences the **sensitivity** of the gamma camera), and the clarity of images (that is the **spatial resolution**) it produces. Unfortunately these two factors are inversely related in that the use of a collimator which produces images of good spatial resolution generally implies that the instrument is not very sensitive to radiation.

Other collimator designs beside the parallel hole type are also in use. For example a **diverging** hole collimator produces a minified image and **converging** hole and **pin-hole** collimators produce a magnified image. The pin-hole collimator is illustrated in the following figure:

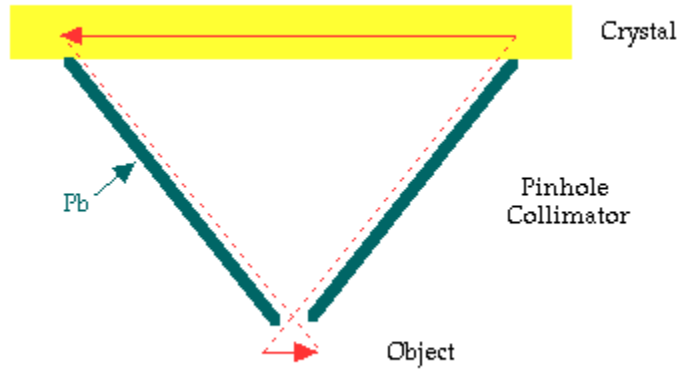


Illustration 3: Diagram of a pin-hole collimator illustrating the inversion of acquired images.

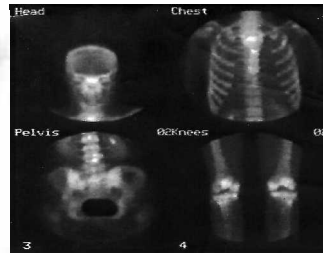
A number of nuclear medicine images are shown below:



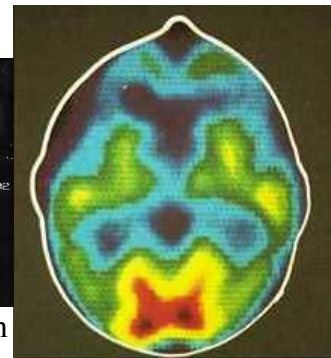
A SPECT slice of the distribution of ^{99m}Tc Ceretec within a patient's brain.



A SPECT slice through a patient's liver.



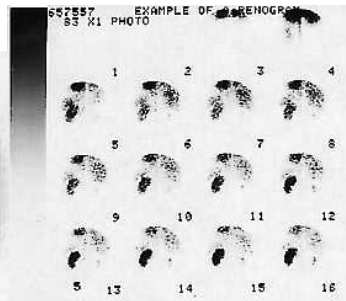
Four planar images from a patient's bone scan (clockwise from top left: head, chest, knees and pelvis).



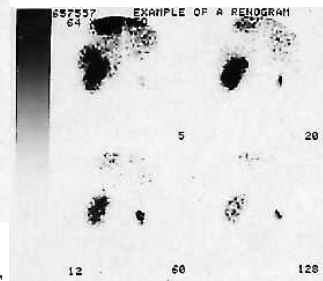
A PET slice of a patient's brain, with a region of interest drawn to indicate the skin surface.



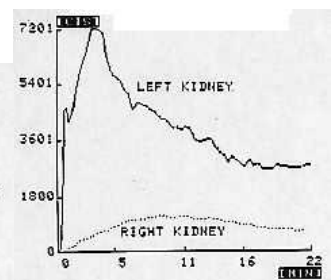
A thyroid scan showing an outline used for computing thyroid uptake.



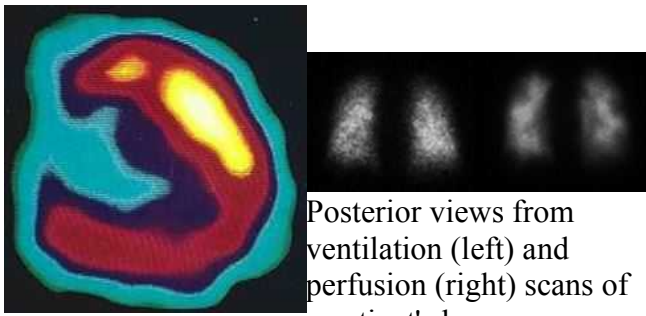
A series of planar images acquired every 10 seconds during a renogram of a patient with a stone blocking their right kidney.



Selected images from a renogram series.



A graphical display showing the number of counts in each kidney versus time for a renogram.



A SPECT slice of a patient's heart. Posterior views from ventilation (left) and perfusion (right) scans of a patient's lungs.

Emission Tomography

The form of imaging which we have been describing is called **Planar Imaging**. It produces a two-dimensional image of a three-dimensional object. As a result images contain no depth information and some details can be superimposed on top of each other and obscured or partially obscured as a result. Note that this is also a feature of conventional X-ray imaging.

The usual way of trying to overcome this limitation is to take at least two views of the patient, one from the front and one from the side for example. So in chest radiography a postero-anterior (PA) and a lateral view can be taken. And in a nuclear medicine liver scan an antero-posterior (AP) and lateral scan are taken.

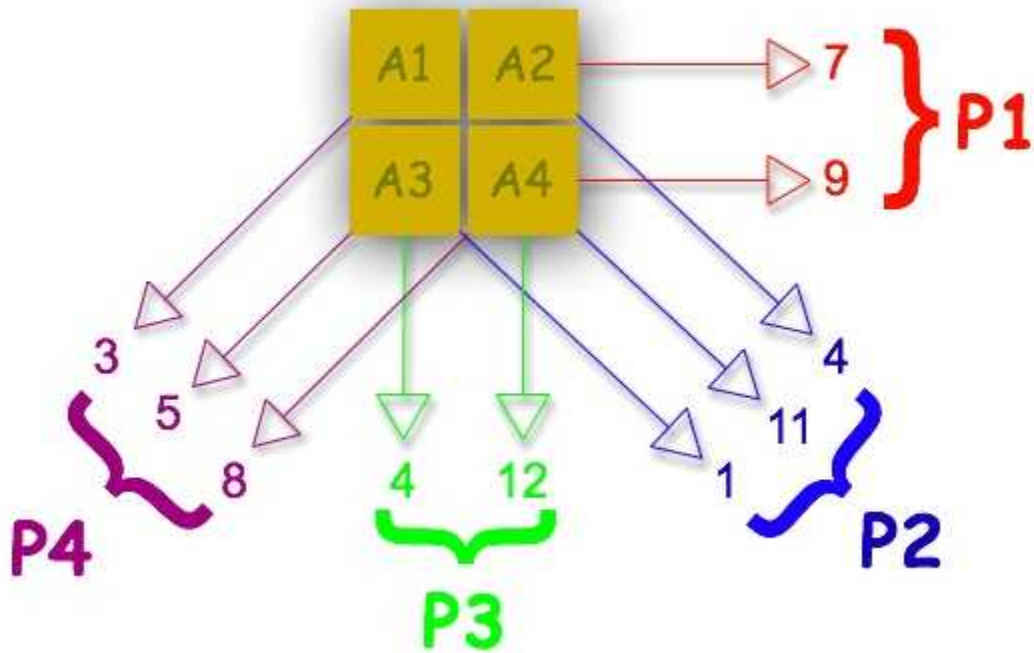
This limitation of planar X-ray imaging was overcome by the development of the **CAT Scanner** about 1970 or thereabouts. CAT stands for Computerised Axial Tomography or Computer Assisted Tomography and today the term is often shortened to Computed Tomography or CT scanning (the term tomography comes from the Greek word *tomos* meaning **slice**). Irrespective of its exact name the technique allows images of slices through the body to be produced using a computer. It does this in essence by taking X-ray images at a number of angles around the patient. These slice images show the third dimension which is missing from planar images and thus eliminate the problem of superimposed details. Furthermore images of a number of successive slices through a region of the patient can be stacked on top of each other using the computer to produce a **three-dimensional image**. Clearly CT scanning is a very powerful imaging technique which is far superior to planar imaging.

The equivalent nuclear medicine imaging technique is called Emission Computed Tomography. We will consider two implementations of this technique below.

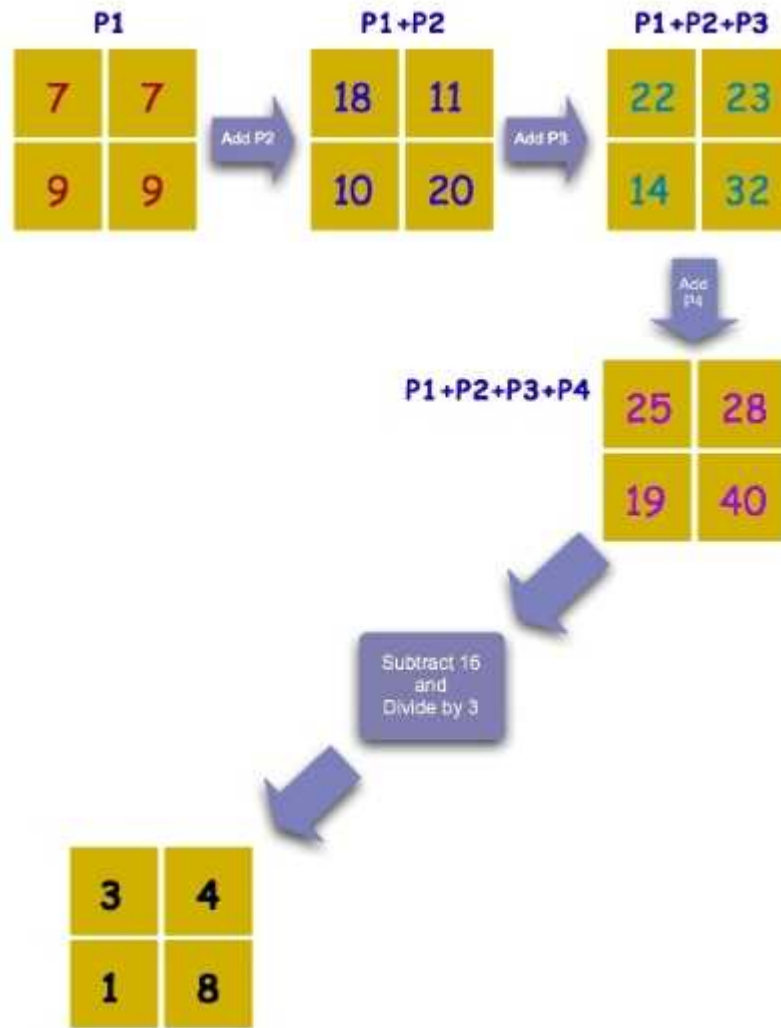
(a) Single Photon Emission Computed Tomography (SPECT)

This **SPECT** technique uses a gamma camera to record images at a series of angles around the patient. These images are then subjected to a form of digital image processing called **Image Reconstruction** in order to compute images of slices through the patient.

The **Filtered Back Projection** reconstruction process is illustrated below. Let us assume for simplicity that the slice through the patient actually consists of a 2x2 voxel array with the radioactivity in each voxel given by $A_1...A_4$:



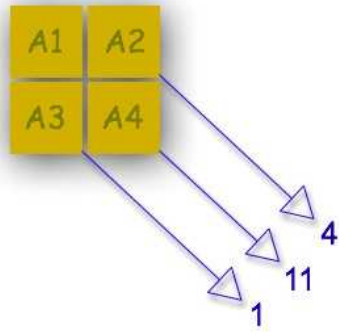
The first projection, P1, is imaged from the right and the second projection, P2, from the right oblique and so on. The projections are firstly added to each as shown below:



and the summed (or superimposed) projections are normalised to generate an estimate of the radioactivity in each voxel.

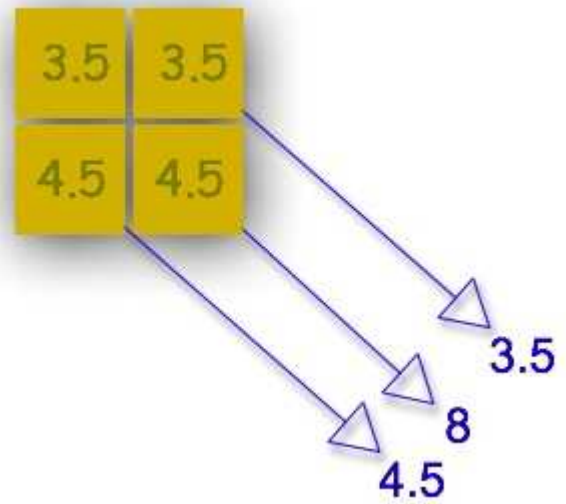
An alternative image reconstruction technique is called **Iterative Reconstruction**. This is a successive approximation technique as illustrated below:

Projection	Patient	Additive Iterative Reconstruction
P1	<p>Actual projection, P1.</p>	<p>First estimate of image matrix.</p>

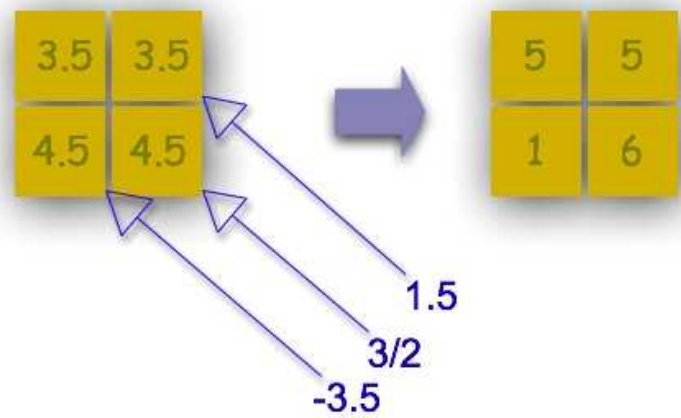


Actual projection, P2.

P2



Estimate of projection, P2.

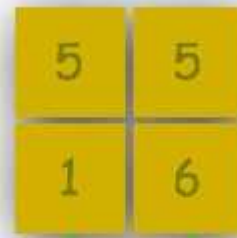


Second estimate of image matrix.

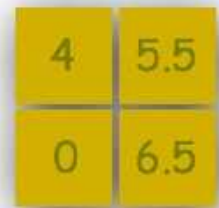
P3



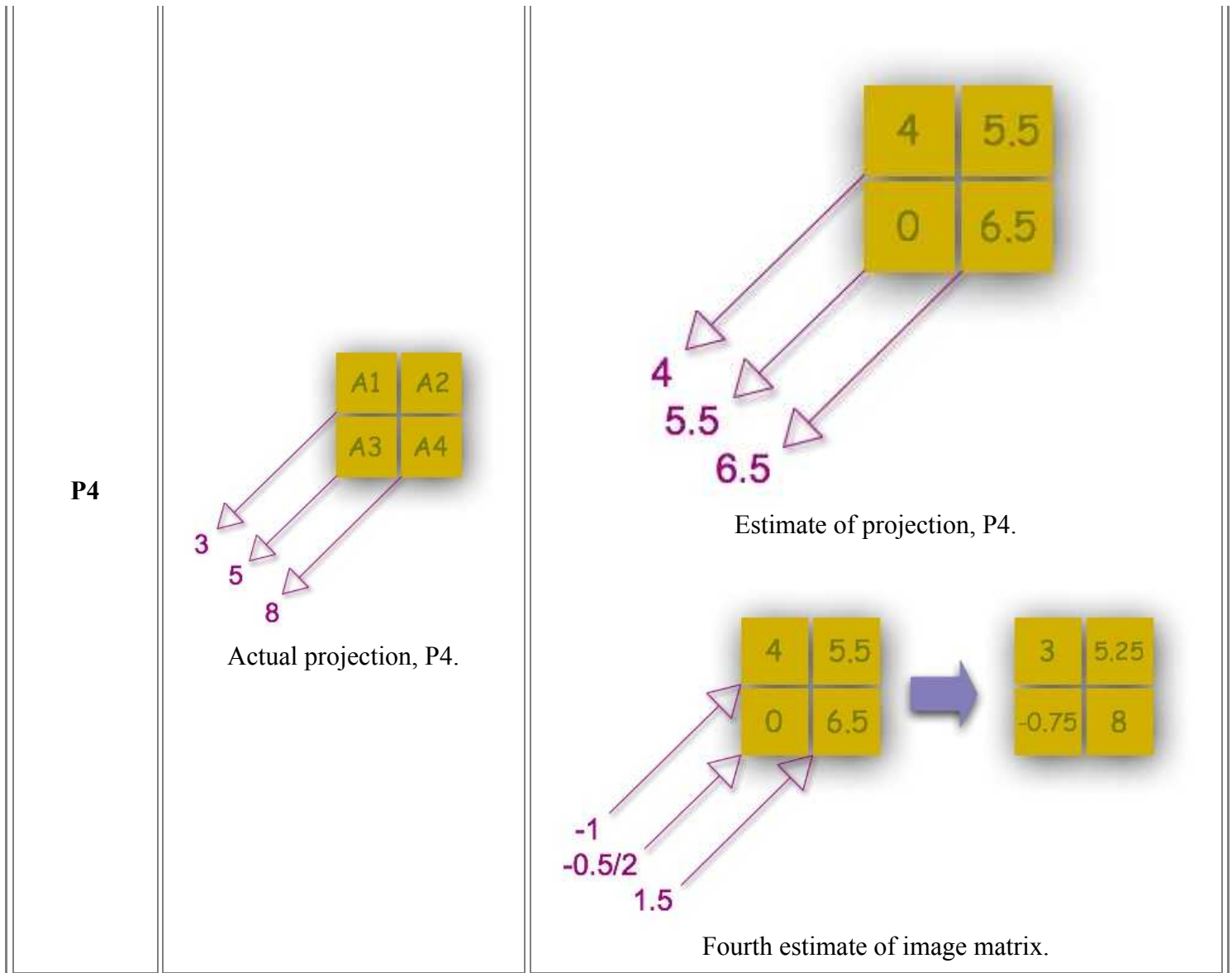
Actual projection, P3.



Estimate of projection, P3.



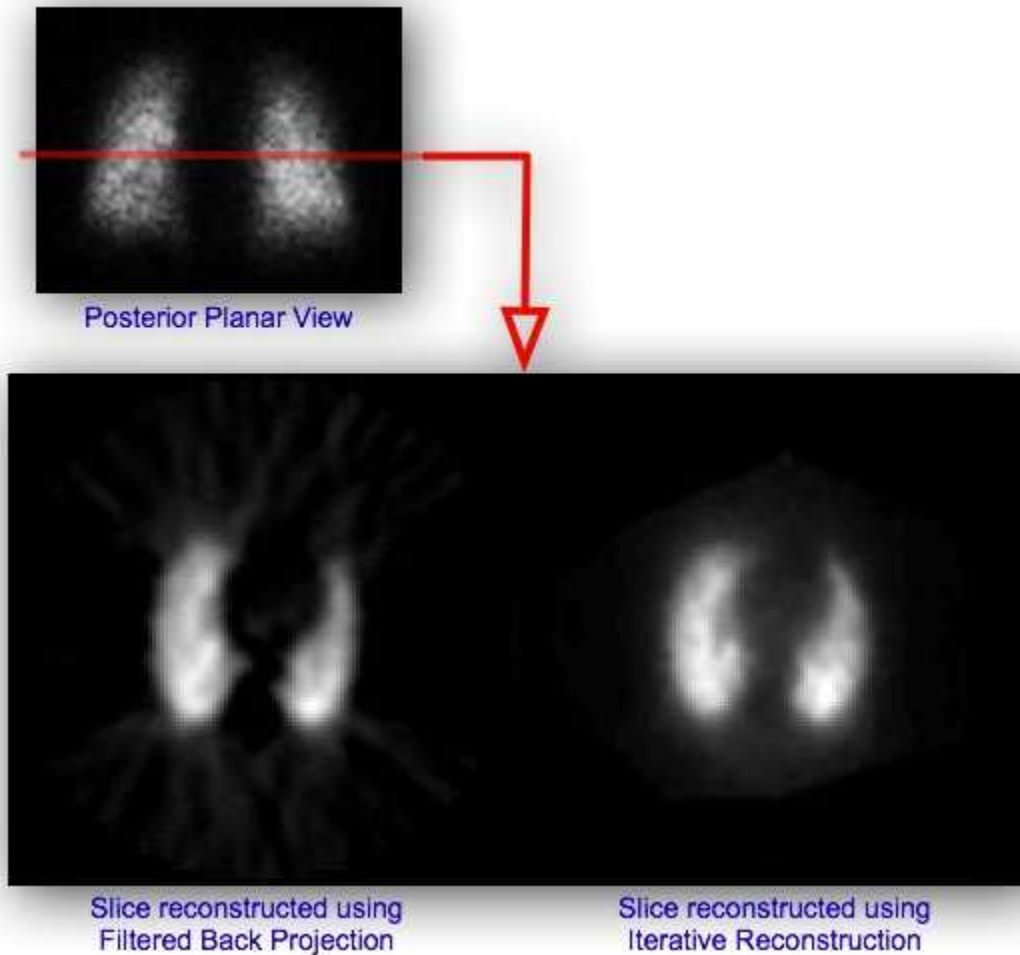
Third estimate of image matrix.



The first estimate of the image matrix is made by distributing the first projection, P1, evenly through an empty pixel matrix. The second projection, P2, is compared to the same projection from the estimated matrix and the difference between actual and estimated projections is added to the estimated matrix. The process is repeated for all other projections.

The gamma camera is typically rotated around the patient in order to acquire the images. Modern gamma cameras which are designed specifically for SPECT scanning can consist of two camera heads mounted parallel to each other with the patient in between. The time required to produce images is therefore reduced by a factor of about two. In addition some SPECT gamma cameras designed for brain scanning have three camera heads mounted in a triangular arrangement.

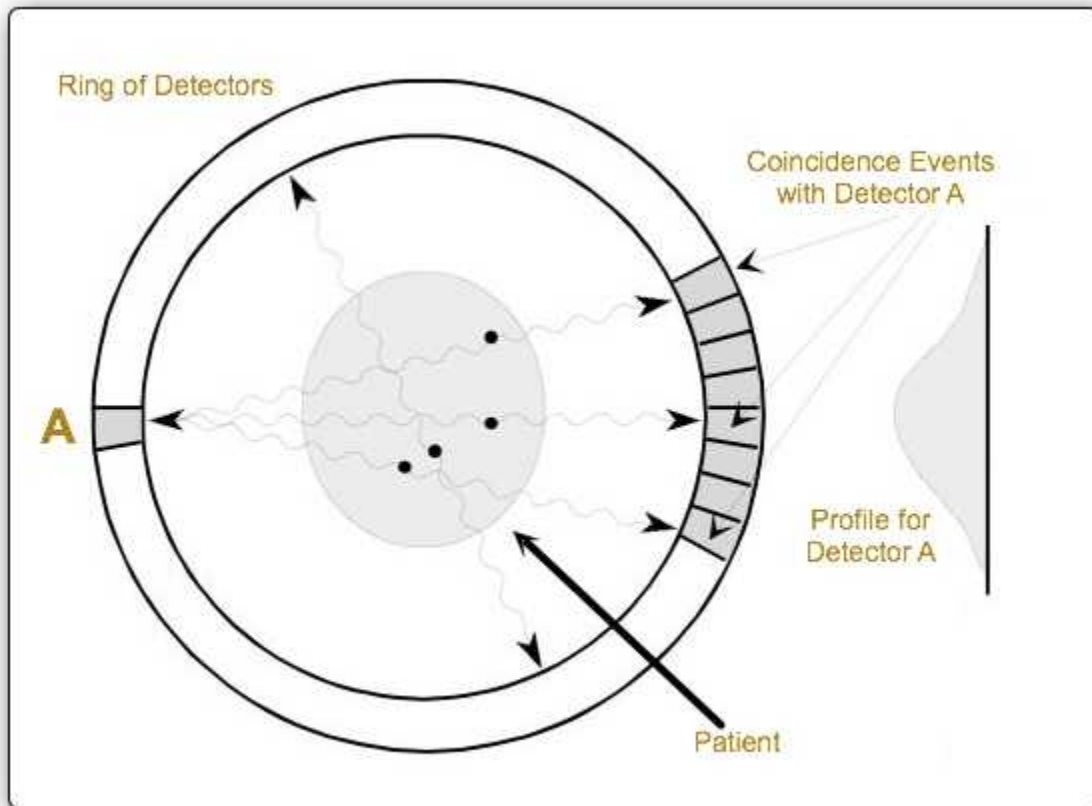
A comparison of these image reconstruction techniques is shown below for a slice through a ventilation scan of a patient's lungs:



(b) Positron Emission Tomography (PET)

You will remember from chapter 2 that positrons can be emitted from radioactive nuclei which have too many neutrons for stability. You will also remember that positrons do not last for very long in matter since they will quickly encounter an electron and a process called **annihilation** results. In the process the positron and electron vanish and their energy is converted into two gamma-rays which are emitted at 180° degrees to each other. The emission is often referred to as two **back-to-back** gamma-rays and they each have a discrete energy of 0.51 MeV.

So if we administer a positron-emitting radiopharmaceutical to a patient an emitted positrons can annihilate with a nearby electron and two gamma-rays will be emitted in opposite directions. These gamma-rays can be detected using a ring of radiation detectors encircling the patient and tomographic images can be generated using a computer system. The detectors are typically specialised scintillation devices which are optimised for detection of the 0.51 MeV gamma-rays. This ring of detectors, associated apparatus and computer system are called a **PET Scanner**:



The locations of positron decays within the patient are highlighted by the solid circles in the above diagram. In addition only a few detectors are shown in the diagram for reasons of clarity. Each detector around the ring is operated in coincidence with a bank of opposing detectors and the annihilation gamma-rays thus detected are used to build up a single profile.

It has also been found that gamma cameras fitted with thick crystals and special collimators can be used for PET scanning.

The radioisotopes used for PET scanning include ^{11}C , ^{13}N , ^{15}O and ^{18}F . These isotopes are usually produced using an instrument called a **cyclotron**. In addition these isotopes have relatively short half lives. PET scanning therefore needs a cyclotron and associated radiopharmaceutical production facilities located close by. We will be considering cyclotrons in the next and **final chapter** of this wikibook.

A number of photographs of a PET scanner are shown below:



circuitry. detectors are under the covering panel.

External Links

- [Centre for Positron Emission Tomography](#) at the Austin & Repatriation Medical Centre, Melbourne with sections on what PET is, current facilities, projects & research and a PET image library.
- [Online Learning Tools](#) - an advanced treatment from the Department of Radiology, Brigham and Women's Hospital, USA containing nuclear medicine teaching files, an atlas of myocardial perfusion SPECT, an atlas of brain perfusion SPECT and the physical characteristics of nuclear medicine images.
- [The Power of Molecular Imaging](#) - a brochure all about Positron Emission Tomography (PET) from the Department of Molecular & Medical Pharmacology, University of California LA, USA.

Multiple Choice Questions

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CHAPTER 10 PRODUCTION OF RADIOISOTOPES

Introduction

Most of the radioisotopes found in nature have relatively long half lives. They also belong to elements which are not handled well by the human body. As a result medical applications generally require the use of radioisotopes which are produced artificially.

We have looked at the subject of radioactivity in earlier chapters of this wikibook and have then progressed to cover the interaction of radiation with matter, radiation detectors and imaging systems. We return to sources of radioactivity in this chapter in order to learn about methods which are used to make radioisotopes.

The type of radioisotope of value to nuclear medicine imaging should have characteristics which keep the radiation dose to the patient as low as possible. For this reason they generally have a short half life and emit only gamma-rays - that is no alpha-particle or beta-particle emissions. From an energy point of view the gamma-ray energy should not be so low that the radiation gets completely absorbed before emerging from the patient's body and not too high that it is difficult to detect. For this reason most of the radioisotopes used emit gamma-rays of medium energy, that is between about 100 and 200 keV. Finally since the radioisotope needs to be incorporated into some form of radiopharmaceutical it should also be capable of being produced in a form which is amenable to chemical, pharmaceutical and sterile processing.

The production methods we will consider are nuclear fission, nuclear bombardment and the radioisotope generator.

Nuclear Fission

We were introduced to spontaneous fission in chapter 2 where we saw that a heavy nucleus can break into a number of fragments. This disintegration process can be induced to occur when certain heavy nuclei absorb neutrons. Following absorption of a neutron such nuclei break into smaller fragments with atomic numbers between about 30 and 65. Some of these new nuclei are of value to nuclear medicine and can be separated from other fission fragments using chemical processes.

The fission process is controlled inside a device called a **nuclear reactor**. One such reactor exists in Australia at Lucas Heights in New South Wales and many others exist throughout the world.

Nuclear Bombardment

In this method of radioisotope production charged particles are accelerated up to very high energies and caused to collide into a target material. Examples of such charged particles are protons, alpha particles and deuterons. New nuclei can be formed when these particles collide with nuclei in the target material. Some of these nuclei are of value to nuclear medicine.

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An example of this method is the production of ^{22}Na where a target of ^{24}Mg is bombarded with deuterons, that is



A deuteron you will remember from chapter 1 is the second most common isotope of hydrogen, that is ^2H . When it collides with a ^{24}Mg nucleus a ^{22}Na nucleus plus an alpha particle is produced. The target is exposed to the deuterons for a period of time and is subsequently processed chemically in order to separate out the ^{22}Na nuclei.

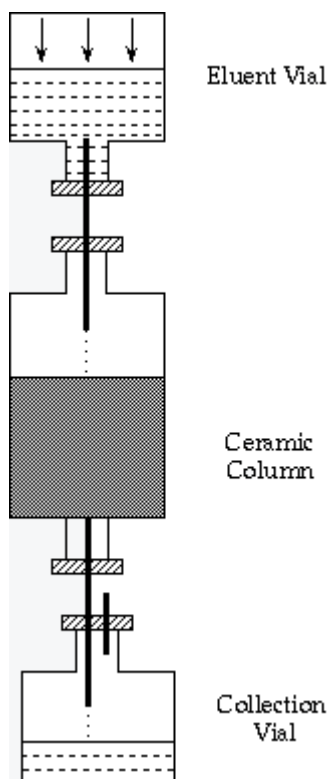
The type of device commonly used for this method of radioisotope production is called a **cyclotron**. It consists of an ion gun for producing the charged particles, electrodes for accelerating them to high energies and a magnet for steering them towards the target material. All arranged in a circular structure.

Radioisotope Generator

This **method** is widely used to produce certain short-lived radioisotopes in a hospital or clinic. It involves obtaining a relatively long-lived radioisotope which decays into the short-lived isotope of interest.

A good example is $^{99\text{m}}\text{Tc}$ which as we have noted before is the most widely used radioisotope in nuclear medicine today. This isotope has a half-life of six hours which is rather short if we wish to have it delivered directly from a nuclear facility. Instead the nuclear facility supplies the isotope ^{99}Mo which decays into $^{99\text{m}}\text{Tc}$ with a half life of about 2.75 days. The ^{99}Mo is called the **parent** isotope and $^{99\text{m}}\text{Tc}$ is called the **daughter** isotope.

So the nuclear facility produces the parent isotope which decays relatively slowly into the daughter isotope and the daughter is separated chemically from the parent at the hospital/clinic. The chemical separation device is called, in this example, a **$^{99\text{m}}\text{Tc}$ Generator**:



It consists of a ceramic column with ^{99}Mo adsorbed onto its top surface. A solution called an **eluent** is passed through the column, reacts chemically with any $^{99\text{m}}\text{Tc}$ and emerges in a chemical form which is suitable for combining with a pharmaceutical to produce a radiopharmaceutical. The arrangement shown in the figure above is called a **Positive Pressure** system where the eluent is forced through the ceramic column by a pressure, slightly above atmospheric pressure, in the eluent vial.

The ceramic column and collection vials need to be surrounded by lead shielding for radiation protection purposes. In addition all components are produced and need to be maintained in a sterile condition since the collected solution will be administered to patients.

Finally an **Isotope Calibrator** is needed when a $^{99\text{m}}\text{Tc}$ Generator is used to determine the radioactivity for preparation of patient doses and to check whether any ^{99}Mo is present in the collected solution.

A number of photographs taken in nuclear medicine hot labs are shown below:



Five 99m-Tc generators.



Pb shielding surrounding the laminar flow cabinet, where radiopharmaceuticals are prepared.



Another view of the Pb shielding surrounding a laminar flow cabinet.

External Links

- [Cyclotron Java Applet](#) - a Java-based interactive demonstration of the operation of a cyclotron from GFu-Kwun Hwang, Dept. of Physics, National Taiwan Normal University, Virtual Physics Laboratory.
- [Nuclear Power Plant Demonstration](#) - a Java-based interactive demonstration of controlling a nuclear reactor. Also contains nuclear power information links.
- [ANSTO's HIFAR Reactor](#) - details about the nuclear reactor at Lucas Heights, NSW.
- [Medical Valley](#) - contains information on what nuclear medicine is, production of nuclear pharmaceuticals, molybdenum and technetium - from The Netherlands Energy Research Foundation Petten.

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CHAPTER 11 CHAPTER REVIEW

Review of Atomic & Nuclear Structure

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- The atom consists of two components - a nucleus (positively charged) and an electron cloud (negatively charged);
- The radius of the nucleus is about 10,000 times smaller than that of the atom;
- The nucleus can have two component particles - **neutrons** (no charge) and **protons** (positively charged) - collectively called **nucleons**;
 - The mass of a proton is about equal to that of a neutron - and is about 1,840 times that of an electron;
 - The number of protons equals the number of electrons in an isolated atom;
 - The **Atomic Number** specifies the number of protons in a nucleus;
 - The **Mass Number** specifies the number of nucleons in a nucleus;
 - **Isotopes** of elements have the same atomic number but different mass numbers;
 - Isotopes are classified by specifying the element's chemical symbol preceded by a superscript giving the mass number and a subscript giving the atomic number;
 - The **atomic mass unit** is defined as 1/12th the mass of the stable, most commonly occurring isotope of carbon (i.e. C-12);
 - **Binding energy** is the energy which holds the nucleons together in a nucleus and is measured in **electron volts** (eV);
 - To combat the effect of the increase in electrostatic repulsion as the number of protons increases, the number of neutrons increases more rapidly - giving rise to the **Nuclear Stability Curve**;
 - There are ~2450 isotopes of ~100 elements and the unstable isotopes lie above or below the Nuclear Stability Curve;
 - Unstable isotopes attempt to reach the stability curve by splitting into fragments (**fission**) or by emitting particles/energy (**radioactivity**);
 - Unstable isotopes \Leftrightarrow radioactive isotopes \Leftrightarrow radioisotopes \Leftrightarrow radionuclides;
 - ~300 of the ~2450 isotopes are found in nature - the rest are produced artificially.

Review of Radioactive Decay

- **Fission**: Some heavy nuclei decay by splitting into 2 or 3 fragments plus some neutrons. These fragments form new nuclei which are usually radioactive;
 - **Alpha Decay**: Two protons and two neutrons leave the nucleus together in an assembly known as an **alpha-particle**;
 - An alpha-particle is a He-4 nucleus;
 - **Beta Decay - Electron Emission**: Certain nuclei with an excess of neutrons may reach stability by converting a neutron into a proton with the emission of a **beta-minus particle**;
 - A beta-minus particle is an electron;
 - **Beta Decay - Positron Emission**: When the number of protons in a nucleus is in excess, the nucleus may reach stability by converting a proton into a neutron with the emission of a **beta-plus particle**;

- A beta-plus particle is a **positron**;
- Positrons **annihilate** with electrons to produce two back-to-back gamma-rays;
- **Beta Decay - Electron Capture**: An inner orbital electron is attracted into the nucleus where it combines with a proton to form a neutron;
- Electron capture is also known as **K-capture**;
- Following electron capture, the excited nucleus may give off some gamma-rays. In addition, as the vacant electron site is filled, an X-ray is emitted;
- **Gamma Decay - Isomeric Transition**: A nucleus in an excited state may reach its ground state by the emission of a gamma-ray;
- A gamma-ray is an electromagnetic photon of high energy;
- **Gamma Decay - Internal Conversion**: the excitation energy of an excited nucleus is given to an atomic electron.

Review of the Radioactive Decay Law

- The radioactive decay law in equation form;
- Radioactivity is the number of radioactive decays per unit time;
- The **decay constant** is defined as the fraction of the initial number of radioactive nuclei which decay in unit time;
- **Half Life**: The time taken for the number of radioactive nuclei in the sample to reduce by a factor of two;
- Half Life = $(0.693)/(\text{Decay Constant})$;
- The SI Unit of radioactivity is the **becquerel** (Bq)
- 1 Bq = one radioactive decay per second;
- The traditional unit of radioactivity is the **curie** (Ci);
- 1 Ci = 3.7×10^{10} radioactive decays per second.

Review of Units of Radiation Measurement

- **Exposure** expresses the intensity of an X- or gamma-ray beam;
- The SI unit of exposure is the **coulomb per kilogram** (C/kg);
- 1 C/kg = The quantity of X- or gamma-rays such that the associated electrons emitted per kg of air at STP produce in air ions carrying 1 coulomb of electric charge;
- The traditional unit of exposure is the **roentgen** (R);
- 1 R = The quantity of X- or gamma-rays such that the associated electrons emitted per kg of air at STP produce in air ions carrying 2.58×10^{-4} coulombs of electric charge;
- The **exposure rate** is the exposure per unit time, e.g. C/kg/s;
- **Absorbed dose** is the radiation energy absorbed per unit mass of absorbing material;
- The SI unit of absorbed dose is the **gray** (Gy);
- 1 Gy = The absorption of 1 joule of radiation energy per kilogram of material;
- The traditional unit of absorbed dose is the **rad**;
- 1 rad = The absorption of 10^{-2} joules of radiation energy per kilogram of material;
- The **Specific Gamma-Ray Constant** expresses the exposure rate produced by the gamma-rays from a radioisotope;
- The Specific Gamma-Ray Constant is expressed in SI units in C/kg/s/Bq at 1 m;
- Exposure from an X- or gamma-ray source follows the **Inverse Square Law** and decreases

with the square of the distance from the source.

Review of the Interaction of Radiation with Matter

- **Alpha-Particles:**
 - exert considerable electrostatic attraction on the outer orbital electrons of atoms near which they pass and cause ionisations;
 - travel in straight lines - except for rare direct collisions with nuclei of atoms in their path;
 - energy is always discrete.
- **Beta-Minus Particles:**
 - attracted by nuclei and repelled by electron clouds as they pass through matter and cause ionisations;
 - have a tortuous path;
 - have a range of energies;
 - range of energies results because two particles are emitted - a beta-particle and a **neutrino**.
- **Gamma-Rays:**
 - energy is always discrete;
 - have many modes of interaction with matter;
 - important interactions for nuclear medicine imaging (and radiography) are the Photoelectric Effect and the Compton Effect.
- **Photoelectric Effect:**
 - when a gamma-ray collides with an orbital electron, it may transfer all its energy to the electron and cease to exist;
 - the electron can leave the atom with a kinetic energy equal to the energy of the gamma-ray less the orbital binding energy;
 - a positive ion is formed when the electron leaves the atom;
 - the electron is called a **photoelectron**;
 - the photoelectron can cause further ionisations;
 - subsequent X-ray emission as the orbital vacancy is filled.
- **Compton Effect:**
 - A gamma-ray may transfer only part of its energy to a valence electron which is essentially free; ** gives rise to a scattered gamma-ray;
 - is sometimes called Compton Scatter;
 - a positive ion results;
- **Attenuation** is term used to describe both absorption and scattering of radiation.

Review of Attenuation of Gamma-Rays

- Attenuation of a narrow-beam of gamma-rays increases as the thickness, the density and the atomic number of the absorber increases;
- Attenuation of a narrow-beam of gamma-rays decreases as the energy of the gamma-rays increases;
- Attenuation of a narrow beam is described by an equation;
- the **Linear Attenuation Coefficient** is defined as the fraction of the incident intensity

absorbed in a unit distance of the absorber;

- Linear attenuation coefficients are usually expressed in units of cm^{-1} ;
- the **Half Value Layer** is the thickness of absorber required to reduce the intensity of a radiation beam by a factor of 2;
- Half Value Layer = $(0.693)/(\text{Linear Attenuation Coefficient})$;
- the **Mass Attenuation Coefficient** is given by the linear attenuation coefficient divided by the density of the absorber;
- Mass attenuation coefficients are usually expressed in units of $\text{cm}^2 \text{g}^{-1}$.

Review of Gas-Filled Detectors

- Gas-filled detectors include the **ionisation chamber**, the **proportional counter** and the **Geiger counter**;
- They operate on the basis of ionisation of gas atoms by the incident radiation, where the positive ions and electrons produced are collected by electrodes;
- An **ion pair** is the term used to describe a positive ion and an electron;
- The operation of gas-filled detectors is critically dependent on the magnitude of the applied dc voltage;
- The output voltage of an ionisation chamber can be calculated on the basis of the capacitance of the chamber;
- A very sensitive amplifier is required to measure voltage pulses produced by an ionisation chamber;
- The gas in ionisation chambers is usually air;
- Ionisation chambers are typically used to measure radiation exposure (in a device called an **Exposure Meter**) and radioactivity (in a device called an **Isotope Calibrator**);
- The total charge collected in a proportional counter may be up to 1000 times the charge produced initially by the radiation;
- The initial ionisation triggers a complete gas breakdown in a Geiger counter;
- The gas in a Geiger counter is usually an inert gas;
- The gas breakdown must be stopped in order to prepare the Geiger counter for a new event by a process called **quenching**;
- Two types of quenching are possible: electronic quenching and the use of a **quenching gas**;
- Geiger counters suffer from **dead time**, a small period of time following the gas breakdown when the counter is inoperative;
- The true count rate can be determined from the actual count rate and the dead time using an equation;
- The value of the applied dc voltage in a Geiger counter is critical, but high stability is not required.

Review of Scintillation Detectors

- NaI(Tl) is a scintillation crystal widely used in nuclear medicine;
- The crystal is coupled to a **photomultiplier tube** to generate a voltage pulse representing the energy deposited in the crystal by the radiation;
- A very sensitive amplifier is needed to measure such voltage pulses;
- The voltages pulses range in amplitude depending on how the radiation interacts with the

crystal, i.e. the pulses form a spectrum whose shape depends on the interaction mechanisms involved, e.g. for medium-energy gamma-rays used in in-vivo nuclear medicine: the Compton effect and the Photoelectric effect;

- A **Gamma-Ray Energy Spectrum** for a medium-energy, monoenergetic gamma-ray emitter consists (simply) of a Compton Smear and a Photopeak;
- **Pulse Height Analysis** is used to discriminate the amplitude of voltage pulses;
- A pulse height analyser (PHA) consists of a **lower level discriminator** (which passes voltage pulses which are than its setting) and an **upper level discriminator** (which passes voltage pulses lower than its setting);
- The result is a variable width window which can be placed anywhere along a spectrum, or used to scan a spectrum;
- A **single channel analyser** (SCA) consists of a single PHA with a scaler and a ratemeter;
- A **multi-channel analyser** (MCA) is a computer-controlled device which can acquire data from many windows simultaneously.

Review of Nuclear Medicine Imaging Systems

- A **gamma camera** consists of a large diameter (25-40 cm) NaI(Tl) crystal, ~1 cm thick;
- The crystal is viewed by an array of 37-91 PM tubes;
- PM tubes signals are processed by a position circuit which generates +/- X and +/- Y signals;
- These position signals are summed to form a **Z signal** which is fed to a pulse height analyser;
- The +/- X, +/- Y and discriminated Z signals are sent to a computer for digital image processing;
- A **collimator** is used to improve the spatial resolution of a gamma-camera;
- Collimators typically consist of a Pb plate containing a large number of small holes;
- The most common type is a **parallel** multi-hole collimator;
- The most resolvable area is directly in front of a collimator;
- Parallel-hole collimators vary in terms of the number of holes, the hole diameter, the length of each hole and the septum thickness - the combination of which affect the sensitivity and spatial resolution of the imaging system;
- Other types include the **diverging**-hole collimator (which generates minified images), the **converging**-hole collimator (which generates magnified images) and the **pin-hole** collimator (which generates magnified inverted images);
- Conventional imaging with a gamma camera is referred to as **Planar Imaging**, i.e. a 2D image portraying a 3D object giving superimposed details and no depth information;
- **Single Photon Emission Computed Tomography** (SPECT) produces images of slices through the body;
- SPECT uses a gamma camera to record images at a series of angles around the patient;
- The resultant data can be processed using **Filtered Back Projection** and **Iterative Reconstruction**;
- SPECT gamma-cameras can have one, two or three camera heads;
- **Positron Emission Tomography** (PET) also produces images of slices through the body;
- PET exploits the positron annihilation process where two 0.51 MeV back-to-back gamma-rays are produced;
- If these gamma-rays are detected, their origin will lie on a line joining two of the detectors of the ring of detectors which encircles the patient;

- A **Time-of-Flight** method can be used to localise their origin;
- PET systems require on-site or nearby **cyclotron** to produce short-lived radioisotopes, such as C-11, N-13, O-15 and F-18.

Review of the Production of Radioisotopes

- Naturally-occurring radioisotopes generally have long half lives and belong to relatively heavy elements - and are therefore unsuitable for medical diagnostic applications;
- Medical diagnostic radioisotopes are generally produced artificially;
- The **fission** process can be exploited so that radioisotopes of interest can be separated chemically from fission products;
 - A **cyclotron** can be used to accelerate charged particles up to high energies so that they to collide into a target of the material to be activated;
 - A radioisotope generator is generally used in hospitals to produce short-lived radioisotopes;
 - A **technetium-99m generator** consists of an alumina column containing Mo-99, which decays into Tc-99m;
 - Saline is passed through the generator to elute the Tc-99m - the resulting solution is called **sodium pertechnetate**;
 - Both positive pressure and negative pressure generators are in use;
 - An **isotope calibrator** is needed when a Tc-99m generator is used in order to determine the activity for preparation of patient doses and to test whether any Mo-99 is present in the collected solution.

Exercise Questions

1. Discuss the process of radioactive decay from the perspective of the nuclear stability curve.
2. Describe in detail FOUR common forms of radioactive decay.
3. Give the equation which expresses the Radioactive Decay Law, and explain the meaning of each of its terms.
4. Define each of the following:
 - Half life
 - Decay Constant
 - Becquerel.
5. A sample of radioactive substance is found to have an activity of 100 kBq. Its radioactivity is measured again 82 days later and is found to be 15 kBq. Calculate:
 - the half-life;
 - the decay constant.
6. Define each of the following radiation units:
 - Roentgen;
 - Becquerel;
 - Gray;
 - Seivert.
7. Estimate the exposure rate at 1 metre from a 100 MBq source of radioactivity which has a Specific Gamma Ray Constant of 50 mR per hour per MBq at 1 cm.
8. Briefly describe the basic principle of operation of gas-filled radiation detectors.
9. Illustrate using a graph how the magnitude of the voltage pulses from a gas-filled radiation

detector varies with applied voltage and identify on the graph the regions associated with the operation of Ionisation Chambers and the Geiger Counters.

10. Describe the construction and principles of operation of a scintillation spectrometer.

11. Discuss the components of the energy spectrum from a monoenergetic, medium energy gamma- emitting radioisotope obtained using a scintillation spectrometer on the basis of how the gamma-rays interact with the scintillation crystal.

12. Describe the construction and principles of operation of a Gamma Camera.

13. Compare features of three types of collimator which can be used with a Gamma Camera.

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